

Gas Laws Practice Worksheet and Study Guide

Constants and Formulas (assume all variables not shown are constant)

Unit Conversions

$$1 \text{ atm} = 101.3 \text{ kPa}$$

$$K = ^\circ\text{C} + 273$$

Constants

$$R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$$

Charles' Law

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

Gay-Lussac's Law

$$\frac{P_1}{T_1} = \frac{P_2}{T_2}$$

Combined Gas Law

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

Boyle's Law

$$P_1 V_1 = P_2 V_2$$

STP

Standard T = 0°C = 273K

Standard P = 1 atm

Molar

volume:

22.4 L/mol at STP

Dalton's Law of Partial Pressure

$$P_{\text{total}} = P_1 + \dots + P_n + P_{\text{water}}$$

Answer on your own paper. Show your work. Circle final answers. Remember, temperatures must be in Kelvin and units must agree to cancel.

Gas Law Practice

Boyle's Law

1. A gas has a volume of 250 mL at 0.29 atm. What will its volume be if the pressure is changed to 0.70 atm?
2. A gas has a volume of .55 L at .83 atm. What will the volume be at a pressure of 1.2 atm?
3. A gas has a volume of 8 liters at 4 atm. What pressure is required to expand the volume to 5500 ml?

Charles's Law

4. A gas has a volume of 6.0 liters at 313 K. What will its volume be at 333K?
5. A gas has a volume of 250 mL at 40°C. If the volume changes to 350 mL, what is the new temperature? (show in °C)

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Rose Army

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