

2.1

CHEMICAL REACTIONS

Section Review

Objectives

1. Measure enthalpy using a calorimeter.
2. Identify and predict chemical reactions using their signs.
3. Apply the law of conservation of energy to chemical reactions.

Vocabulary

- | | | |
|----------------------|---------------|----------------------------|
| 1. endothermic | 2. exothermic | 3. enthalpy |
| 4. activation energy | 5. product | 6. thermochemical equation |

Part A Completion

Write the coefficient needed to balance each chemical equation. Balance each equation by adding whole numbers in front of each formula.

- | | | | |
|--|---|---|---|
| 1. Aluminum reacts with oxygen to form aluminum oxide. | 1 | 2 | 3 |
| 2. Carbon monoxide reacts with oxygen to form carbon dioxide. | 2 | 2 | 1 |
| 3. Iron reacts with oxygen to form iron(III) oxide. | 4 | 3 | 2 |
| 4. Hydrogen reacts with oxygen to form water. | 2 | 1 | 2 |
| 5. The decomposition of hydrogen peroxide into water and oxygen. | 2 | 2 | 1 |
| 6. The decomposition of potassium chlorate into potassium chloride and oxygen. | 2 | 2 | 3 |
| 7. The reaction of aluminum with hydrochloric acid to form aluminum chloride and hydrogen gas. | 2 | 6 | 3 |
| 8. The reaction of sodium metal with water to form sodium hydroxide and hydrogen gas. | 2 | 2 | 1 |

Part B True-False

Write true or false. Explain any false statements.

1. The law of conservation of mass states that matter is neither created nor destroyed in a chemical reaction.
2. The law of conservation of energy states that energy is neither created nor destroyed in a chemical reaction.
3. The law of conservation of mass states that matter is neither created nor destroyed in a chemical reaction.
4. The law of conservation of energy states that energy is neither created nor destroyed in a chemical reaction.

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CH Cherryholmes



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