

Liam Francis, Ethan Priore
Period 3
September 22, 2021

Aluminum foil thickness lab

Purpose: To determine the average thickness of a piece of aluminum foil

Procedure:

1. Mass one piece of aluminum foil.
2. Note the mass and measure its length and width
3. Repeat steps one and two for the second and third pieces of aluminum foil.
4. Use the actual density of aluminum foil to find each pieces volume
5. Divide each volume by the length times the width to find the height

Data:

Aluminum foil density: 2.71 g/cm^3

	Mass	Length	Width	Volume(base d on density)	Thickness (volume/lw)
Piece 1	0.16g	6.8cm	6.0cm	0.059 cm^3	0.0014cm
Piece 2	0.18g	6.5cm	6.9cm	0.066 cm^3	0.0015cm
Piece 3	0.18g	8.4cm	4.9cm	0.066 cm^3	0.0016cm

Average of Thicknesses: 0.0015cm

Analys:

Based on our mass, length, width, and the density of alumanum foil, we can find the volume by dividing our mass by the density. Since volume is defined as length * width * height, we can find the height by dividing the volume by length * width.

Piece 1:

$$(0.16\text{g})/(2.71\text{g/cm}^3) = 0.059\text{cm}^3 \quad (0.059\text{cm}^3)/(6.8*6.0) = 0.0014 \text{ cm}$$

Piece 2:

$$(0.18\text{g})/(2.71\text{g/cm}^3) = 0.066\text{cm}^3 \quad 0.066/(6.5*6.9) = 0.0015 \text{ cm}$$

Piece 3:

$$(0.18\text{g})/(2.71\text{g/cm}^3) = 0.066\text{cm}^3 \quad 0.059/(8.4*4.9) = 0.0016 \text{ cm}$$

Aluminum Foil Thickness Lab Answers

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Aluminum Foil Thickness Lab Answers:

Reviewing **Aluminum Foil Thickness Lab Answers**: Unlocking the Spellbinding Force of Linguistics

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