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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

bond	H-H	O-H
bond energy in kJ/mol	436	463

Use the bond energies in Table 3.1 and ΔH to calculate the bond energy of an O-H bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.

$$(4H_2) + 3(O_2)$$

$$4(2H-H) + 3(2O=O)$$

$$8(436) + 3(498) = 3150 \text{ kJ}$$
- Calculate the energy released when bonds form in the products.

$$\text{bond broken} - \text{bond formed} = -90$$

$$3150 - x = -90$$

$$x = 3150 + 90$$

$$x = 3240 \text{ kJ}$$
- Calculate the energy of per O-H bond.

$$6 \times \text{O-H} = 3240$$

$$\text{O-H} = \frac{3240}{6} = 540 \text{ kJ/mol}$$

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Yijin Wang



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Table of Contents 0620 1 Chem May Jun 02 Egypt Ig Student Room

1. Understanding the eBook 0620 1 Chem May Jun 02 Egypt Ig Student Room
 - The Rise of Digital Reading 0620 1 Chem May Jun 02 Egypt Ig Student Room
 - Advantages of eBooks Over Traditional Books
2. Identifying 0620 1 Chem May Jun 02 Egypt Ig Student Room
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 0620 1 Chem May Jun 02 Egypt Ig Student Room
 - User-Friendly Interface
4. Exploring eBook Recommendations from 0620 1 Chem May Jun 02 Egypt Ig Student Room
 - Personalized Recommendations
 - 0620 1 Chem May Jun 02 Egypt Ig Student Room User Reviews and Ratings
 - 0620 1 Chem May Jun 02 Egypt Ig Student Room and Bestseller Lists

5. Accessing 0620 1 Chem May Jun 02 Egypt Ig Student Room Free and Paid eBooks
 - 0620 1 Chem May Jun 02 Egypt Ig Student Room Public Domain eBooks
 - 0620 1 Chem May Jun 02 Egypt Ig Student Room eBook Subscription Services
 - 0620 1 Chem May Jun 02 Egypt Ig Student Room Budget-Friendly Options
6. Navigating 0620 1 Chem May Jun 02 Egypt Ig Student Room eBook Formats
 - ePub, PDF, MOBI, and More
 - 0620 1 Chem May Jun 02 Egypt Ig Student Room Compatibility with Devices
 - 0620 1 Chem May Jun 02 Egypt Ig Student Room Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 0620 1 Chem May Jun 02 Egypt Ig Student Room
 - Highlighting and Note-Taking 0620 1 Chem May Jun 02 Egypt Ig Student Room
 - Interactive Elements 0620 1 Chem May Jun 02 Egypt Ig Student Room
8. Staying Engaged with 0620 1 Chem May Jun 02 Egypt Ig Student Room
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 0620 1 Chem May Jun 02 Egypt Ig Student Room
9. Balancing eBooks and Physical Books 0620 1 Chem May Jun 02 Egypt Ig Student Room
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 0620 1 Chem May Jun 02 Egypt Ig Student Room
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 0620 1 Chem May Jun 02 Egypt Ig Student Room
 - Setting Reading Goals 0620 1 Chem May Jun 02 Egypt Ig Student Room
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 0620 1 Chem May Jun 02 Egypt Ig Student Room
 - Fact-Checking eBook Content of 0620 1 Chem May Jun 02 Egypt Ig Student Room
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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