

3. The amount of solute required to form a saturated solution in a particular amount of solvent is the solubility of a substance.
4. The interaction between solvent and solute particles is called solvation.
5. The separation of charge on a water molecule is a(n) dipole.
6. A(n) exothermic reaction gives off energy.

On the line at the left, indicate if the solubility of the solute increases, decreases, or remains constant for each situation.

- | | |
|-------------------------|--|
| <u>decreases</u> | 7. raising the temperature of a liquid solution when the solute is a gas |
| <u>remains constant</u> | 8. raising the temperature of a liquid solution, if the temperature remains constant when a solid solute is mixed with the solvent |
| <u>decreases</u> | 9. raising the temperature of a liquid solution, if the temperature rises when a solid solute is mixed with the solvent |
| <u>increases</u> | 10. raising the temperature of a liquid solution, if the temperature drops when a solid solute is mixed with the solvent |
| <u>increases</u> | 11. increasing the pressure above a liquid solution when the solute is a gas |

Answer each of the following questions in the space provided.

12. As sugar is dissolved in water, explain how a solution is formed.

The water molecules are attracted to the

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Michael Rajnik



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