

## Chemistry 110 Laboratory

### Synthesis of Alum

#### Introduction:

Aluminum recycling is a major industry. In many cases the recovered aluminum, from drink cans or other aluminum scrap, is converted to useful compounds. One such compound made from aluminum is called "alum", and has the formula  $\text{KAl}(\text{SO}_4)_2 \cdot 12 \text{H}_2\text{O}$ . This compound, a white crystalline solid, has diverse uses, ranging from water purification to the making of pickles. It is also known as "kalinite".

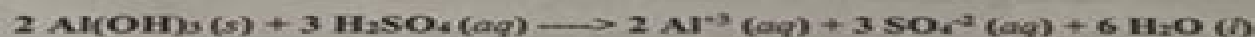
In this lab, alum will be synthesized through the following sequential reactions:



The above reaction dissolves the aluminum metal.



The above reaction converts the potassium aluminum hydroxide to the aluminum hydroxide precipitate.



The above reaction dissolves the aluminum hydroxide precipitate and provides the aqueous ions from which the alum will crystallize, as shown in the reaction below:



#### Procedure:

1. Weigh about 1 gram of aluminum foil to the nearest 0.01 g. Tear the aluminum into small pieces and place them in a clean 250-mL Erlenmeyer flask.
2. Use a 100-mL graduated cylinder to obtain 50 mL of 2 M KOH solution, and add it to the flask. (Handle the KOH solution with care. It is corrosive.)

# Alum From Aluminum Lab Answers

**Lingjun Ying**



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