

2. For Iridium-178 (atomic number 77)

a. Determine the number of protons, neutrons, and electrons

#p = 77

#n = 101

#e = 77

b. Write its electron configuration using noble gas notation



c. How many unpaired electrons does it contain? 3

33. Name the element represented by each of the following electron configurations and determine the number of unpaired electrons each contains:

a. $1s^22s^22p^63s^23p^4$

Sulfur

unpaired e⁻

2

b. $[Kr]5s^24d^8$

Rhodium

2

c. $[Xe]6s^24f^{14}5d^{10}6p^2$

Lead

2

34. Determine which of the following electron configurations represent an atom in the excited state and identify that element. (Circle choice and name the element)

$\uparrow$

a. $1s^22s^32p^63s^23p^6$

b. $1s^22s^12p^63s^23p^5$
Sulfur

c. $1s^23s^22s^22p^63p^6$

35. Three isotopes of argon occur in nature, Ar-36, Ar-38, and Ar-40. Calculate the atomic mass of argon to two decimal places, given the following relative atomic masses and abundances of each of the isotopes: argon-36 (35.97 u; 0.337%), argon-38 (37.96 u; 0.063%), and argon-40 (39.96 u; 99.600%).

$$0.00337(35.97) + 0.00063(37.96) + 0.99600(39.96) = \boxed{39.95 \text{ u}}$$

36. Calculate the mass in grams of each of the following:

a. 3.00 mol Al

$$3.00 \text{ mol Al} \times \frac{26.98 \text{ g}}{1 \text{ mol}} = \boxed{80.9 \text{ g}}$$

b. 2.56×10^{24} atoms Li

$$2.56 \times 10^{24} \text{ atoms} \times \frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ atoms}} \times \frac{6.94 \text{ g}}{1 \text{ mol}} = \boxed{29.5 \text{ g Li}}$$

Calculating Average Atomic Mass Worksheet Answers

**Atomic Energy of Canada
Limited, Alastair Graham Walter
Cameron**



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Holt Chemistry Ralph Thomas Myers, 2004 **Fundamentals of Analytical Chemistry** Douglas A. Skoog, 2004 This text is known for its readability combined with a systematic rigorous approach Extensive coverage of the principles and practices of quantitative chemistry ensures suitability for chemistry majors **How to Calculate Relative Mass, Atomic Mass and Empirical Formula** E Staff, Learn and review on the go Use Quick Review Chemistry Study Notes to help you learn or brush up on the subject quickly You can use the review notes as a reference to understand the subject better and improve your grades Easy to remember facts to help you perform better Learn how to resolve various reaction problems Perfect study notes for all high school health sciences premed medical and nursing students **Atomic and Molecular Weight Determination** Richard Bell Moyes, 1971 **A New Empirical Mass Equation** Harris B. Levy, 1956 Revised Semiempirical Atomic Mass Formula Atomic Energy of Canada Limited, Alastair Graham Walter Cameron, 1959 **Atomic Masses and Fundamental Constants 6** Jerry A. Nolen, 2012-12-06 The Sixth International Conference on Atomic Masses was held in East Lansing Michigan Sept 18 21 1979 The conference was initiated organized and sponsored by the Commission on Atomic Masses and Fundamental Constants of the International Union of Pure and Applied Physics The members of the conference committee are listed below W Benenson Chairman Michigan State University R C Barber University of Manitoba E R Cohen Rockwell International Institute of Chemical Physics V I Goldanskii Moscow J C Hardy Chalk River Canada W H Johnson University of Minnesota E Kashy Michigan State University Orsay France R Klapisch J A Nolen Jr Michigan State University R G H Robertson Michigan State University E Roeckl G S I Darmstadt B N Taylor National Bureau of Standards O Schult IKF Julich A H Wapstra IFO Amsterdam N Zeldes Racah Institute Jerusalem The conference was a little different from the preceding one in Paris 1975 in that the fundamental constant aspects were limited to those directly related to atomic masses The gap is to be filled by the second International Conference on Precision Measurement and Fundamental Constants which is now scheduled for June 1981 in Gaithersburg Maryland Only one of the seven sessions in this conference was devoted to fundamental constant determinations The conference was very strongly supported by the Department of Energy the National Science Foundation and the International Union of Pure and Applied Physics **Atomic Masses and Fundamental Constants 5** J. Sanders, Aaldert Hendrik Wapstra, 1976-03 **ATOMIC STRUCTURE** NARAYAN CHANGDER, 2024-05-01 Note Anyone can request the PDF version of this practice set workbook by emailing me at cbsenet4u gmail com You can also get full PDF books in quiz format on our youtube channel <https://www.youtube.com/smartquiziz> I will send you a PDF version of this workbook This book has been designed for candidates preparing for various competitive examinations It contains many objective questions specifically designed for different exams Answer keys are provided at the end of each page It will undoubtedly serve as the best preparation material for aspirants This book is an engaging quiz eBook for all and offers something for everyone This book will satisfy the curiosity of most

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Atomic Masses and Fundamental Constants 4 J. Sanders, Aaldert Hendrik Wapstra, 1972 The Conference on Nuclear Masses and their Determination which was held at the Max Planck Institute in Mainz from 10 to 12 July 1956 resulted in the formation by the International Union of Pure and Applied Physics of a Commission on Atomic Masses and Related Constants Under the auspices of this Commission conferences covering these subjects were held in Hamilton Ontario 12-16 September 1960 in Vienna 15-19 July 1963 and in Winnipeg 28 August-1 September 1967 After the last of these conferences the Commission felt that the position regarding nuclear masses was reasonably good and that as a consequence the fundamental constants should get more emphasis in future conferences For this reason they were very pleased to accept an offer from the National Physical Laboratory to accept the main burden of organizing the Fourth International Conference on Atomic Masses and Fundamental Constants An Organizing Committee was appointed with the members J H Sanders Chairman B W Petley Secretary A Horsfield Treasurer P Dean A H Wapstra They were assisted by a Program Committee consisting of A H Wapstra Chairman E R Cohen A Horsfield W H Johnson Jr J H Sanders J Terrien

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