

EDTA Titration for Determination of Calcium and Magnesium

Before attempting this experiment, you may need to consult the section in the textbook dealing with complexometric titrations. In this procedure a stock solution of EDTA will be standardized against primary standard CaCO_3 . This standardized EDTA solution is then used to determine water "hardness" – the total calcium and magnesium content, typically expressed as equivalent calcium carbonate mass concentration. While both metals are nutrients that are needed for good human health, high concentrations of these cations (Ca^{2+} and Mg^{2+}) in water can cause deposits to form in bathrooms and kitchens, and cause the formation of "soap scum" – clogging pipes and decreasing the effectiveness of the cleanser. Therefore, very hard water is undesirable for use as tap water. Guidelines for classification of water as provided by the U.S. Geologic Survey are listed in the table below. The USGS website has more information about water hardness and alkalinity if you are interested. We're fortunate in that most of the water in the Portland area is soft.

Table 1 U.S. Geological Survey Classification of Water Hardness

(Source: <http://water.usgs.gov/owq/hardness-alkalinity.html>)

CaCO_3 (mg/L)	Hardness
0-60	soft
61-120	moderately hard
121-180	hard
>180	very hard

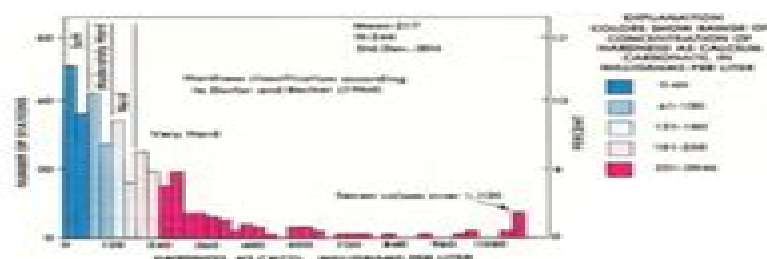


Figure 1 Hardness classification from Briggs et al. 1977

Briggs, J.C., and Ficke, J.F., 1977, Quality of Rivers of the United States, 1975 Water Year -- Based on the National Stream Quality Accounting Network (NASQAN): U.S. Geological Survey Open-File Report 78-200, 436 p.

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Handbook of Water Analysis Leo M.L. Nollet, Leen S. P. De Gelder, 2000-06-27 This work details water sampling and preservation methods by enumerating the different ways to measure physical chemical organoleptical and radiological characteristics It provides step by step descriptions of separation residue determination and cleanup techniques for a variety of fresh and salt waters It also discusses information regarding the analysis and detection of bacteria and algae

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Chemical Analysis Herbert August Laitinen, 1960 Equilibrium and activity Acid base equilibria in water Acid base equilibria in nonaqueous solvents Applications of acid base titrations Solubility of precipitates the formation of precipitates Colloidal properties of precipitates Aging of precipitates Contamination of precipitates Thermal decomposition and volatilization

Selected Technical Publications United States. Food and Drug Administration, 1966 *A Brief Introduction to Quantitative Chemical Analysis* Robert Blanchard Fischer, Dennis G. Peters, 1969 *Inorganic Titrimetric Analysis* Walter Wagner, Clarence Joseph Hull, 1971 Report - US Army Medical Research Laboratory Army Medical Research Laboratory, 1947 *Treatise on Titrimetry: Inorganic titrimetric analysis, by W. Wagner and C. J. Hull* Joseph Jordan, 1971

An Introduction to Chemical Analysis Walter Edgar Harris, Byron Kratochvil, 1981 *Bulletin*, 1977

Quantitative Analysis Harvey Diehl, 1970 **Titrimetric Analysis** Norman B. B. Johnstone, Thomas C. Downie, 1961

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