

Calculations in Analytical Chemistry

The Mole

The mole is the SI unit for amount of substance. It corresponds to the amount of a chemical substance that contains the same number of particles present in 12 grams of C-12 which is equivalent to Avogadro's number, that is, $N_A = 6.022 \times 10^{23}$.

The molar mass is given as the mass of a compound in grams per mole of that substance. It is calculated by adding the molar masses of all the component elements in the compound. Take glucose for example:

Element	Molar mass per element	Number of atoms	Mass
C	12 g/mol	6	72 g/mol
H	1 g/mol	12	12 g/mol
O	16 g/mol	6	96 g/mol
Molar Mass			180 g/mol

Mole and Millimole Calculations

The millimole is simply 1/1000 of a mole. It is sometimes used in calculations because of its convenience.

Example (Skoog, et al., 2014):

1. Find the number of moles and millimoles of benzoic acid (122.1 g/mol) that are contained in 2.00 g of pure acid.

If we use HBz to represent benzoic acid:

$$HBz_{mol} = 2.00 \cancel{g_{HBz}} \times \frac{1 \cancel{mol_{HBz}}}{122.1 \cancel{g_{HBz}}} = 0.0164 \cancel{mol_{HBz}}$$

To obtain the number of millimoles, the millimolar mass of benzoic acid will be used:

$$HBz_{mol} = 2.00 \cancel{g_{HBz}} \times \frac{1 \cancel{mmol_{HBz}}}{0.1221 \cancel{g_{HBz}}} = 16.4 \cancel{mmol_{HBz}}$$

2. What is the mass in grams of Na^+ (22.99 g/mol) in 25.0 g of Na_2SO_4 (142.0 g/mol)?

$$mass_{Na^+} = 25.0 \cancel{g_{Na_2SO_4}} \times \frac{1 \cancel{mol_{Na_2SO_4}}}{142.0 \cancel{g_{Na_2SO_4}}} \times \frac{2 \cancel{mol_{Na^+}}}{1 \cancel{mol_{Na_2SO_4}}} \times \frac{22.99 \cancel{g_{Na^+}}}{1 \cancel{mol_{Na^+}}}$$

$$mass_{Na^+} = 8.10 g_{Na^+}$$

Calculations In Chemistry An Introduction

Alexander Smith



Calculations In Chemistry An Introduction:

Calculations in Chemistry Donald J. Dahm, Eric A. Nelson, 2013 A research backed learning system for preparatory chemistry that systematically reinforces both the fundamental concepts and math skills that are prerequisites for success in general chemistry Recent research has documented that in K 12 programs over the past decade instruction in math computation has been emphasized Calculations in Chemistry addresses these gaps in student background applying the findings of recent scientific research on cognition fluency in applied mathematics and reading comprehension to prepare students for the rigor and pace of general chemistry

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Introduction to Chemistry Calculations B. A. Matthews, K. R. Winter, 1982

Introduction to general inorganic chemistry Alexander Smith, 1906 Introduction to Inorganic Chemistry Alexander Smith, 1906 Principles and Perspectives in Cosmochemistry Aruna Goswami, B. Eswar Reddy, 2010-03-19 A fundamental question in contemporary astrophysics is the origin of the elements Cosmochemistry seeks to answer when how and where the chemical elements arose Quantitative answers to these fundamental questions require a multi disciplinary approach involving stellar evolution explosive nucleosynthesis and nuclear reactions in different astrophysical environments There remain however many outstanding problems and cosmochemistry remains a fertile area of research This book is among the first in recent times to put together the essentials of cosmochemistry combining contributions from leading astrophysicists in the field The chapters have been organized to provide a clear description of the fundamentals an introduction to modern techniques such as computational modelling and glimpses of outstanding issues

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Library, Philadelphia, 1913 **Introduction to Organic Research** Ebenezer Emmet Reid, 1924 *Introduction to Symmetry and Group Theory for Chemists* Arthur M. Lesk, 2007-05-08 This book is based on a one semester course for advanced undergraduates specializing in physical chemistry I am aware that the mathematical training of most science majors is more heavily weighted towards analysis typically calculus and differential equations than towards algebra But it remains my conviction that the basic ideas and applications of group theory are not only vital but not difficult to learn even though a formal mathematical setting with emphasis on rigor and completeness is not the place where most chemists would feel most comfortable in learning them The presentation here is short and limited to those aspects of symmetry and group theory that are directly useful in interpreting molecular structure and spectroscopy Nevertheless I hope that the reader will begin to sense some of the beauty of the subject Symmetry is at the heart of our understanding of the physical laws of nature If a reader is happy with what appears in this book I must count this a success But if the book motivates a reader to move deeper into the subject I shall be gratified **Register of the University of California** University of California (1868-1952), 1955 **Subject Index of the Modern Books Acquired by the British Museum in the Years ...** , 1927

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