

Experiment 10: Electrochemistry with a Standard Solution from: *Basic Laboratory Experiments*, Experiment 10, 10-1, 10-2

The purpose of this experiment is to determine the standard reduction potential for several half-reactions. In this experiment, a standard solution of a metal ion is used to determine the standard reduction potential of a half-reaction. The standard reduction potential of a half-reaction is determined by measuring the potential of a cell in which the half-reaction is coupled to a standard hydrogen electrode. The standard hydrogen electrode is a half-reaction in which hydrogen gas is oxidized to hydrogen ions. The standard reduction potential of a half-reaction is determined by measuring the potential of a cell in which the half-reaction is coupled to a standard hydrogen electrode. The standard reduction potential of a half-reaction is determined by measuring the potential of a cell in which the half-reaction is coupled to a standard hydrogen electrode.

In this experiment, you will determine the standard reduction potential of several half-reactions. The standard reduction potential of a half-reaction is determined by measuring the potential of a cell in which the half-reaction is coupled to a standard hydrogen electrode.

Objectives:

1. to determine the standard reduction potential of several half-reactions
2. to determine the standard potential of the standard hydrogen electrode
3. to determine the standard potential of several half-reactions
4. to determine the standard potential of several half-reactions
5. to determine the standard potential of several half-reactions

Materials:

- standard solution
- standard solution (100 mL)
- standard solution
- standard solution
- standard solution
- standard solution
- standard solution
- standard solution
- standard solution
- standard solution

Procedure:

1. standard solution
2. standard solution
3. standard solution
4. standard solution

Results:

1. standard solution
2. standard solution

Discussion:

1. This is a good laboratory experiment.
2. This experiment is a good laboratory experiment. The standard reduction potential of a half-reaction is determined by measuring the potential of a cell in which the half-reaction is coupled to a standard hydrogen electrode.

Heath Chemistry 11 Lab 6b Answers Warpig

Xiaolong Qi



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