

#### Ideal Gas Law Worksheet PV = nRT

Use the ideal gas law, "PerV-nRT", and the universal gas constant  $R = 0.0821 \frac{L*atm}{L*mol}$  to solve the following problems:

If pressure is needed in kPa then convert by multiplying by 101.3kPa / latm to get

R =8.31 kPa\*L / (K\*mole)

1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of 87 °C, how many moles of gas do I have?

3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

4) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of 56 °C, what is the volume of the container that the gas is in?

5) If I have 17 moles of gas at a temperature of 67 °C, and a volume of 88.89 liters, what is the pressure of the gas?

6) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have?

# **Gas Law Formula Sheet Answers**

## American Association for the Advancement of Science

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#### **Table of Contents Gas Law Formula Sheet Answers**

- 1. Understanding the eBook Gas Law Formula Sheet Answers
  - The Rise of Digital Reading Gas Law Formula Sheet Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Gas Law Formula Sheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gas Law Formula Sheet Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Gas Law Formula Sheet Answers
  - Personalized Recommendations
  - Gas Law Formula Sheet Answers User Reviews and Ratings

- Gas Law Formula Sheet Answers and Bestseller Lists
- 5. Accessing Gas Law Formula Sheet Answers Free and Paid eBooks
  - Gas Law Formula Sheet Answers Public Domain eBooks
  - Gas Law Formula Sheet Answers eBook Subscription Services
  - Gas Law Formula Sheet Answers Budget-Friendly Options
- 6. Navigating Gas Law Formula Sheet Answers eBook Formats
  - o ePub, PDF, MOBI, and More
  - Gas Law Formula Sheet Answers Compatibility with Devices
  - Gas Law Formula Sheet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Gas Law Formula Sheet Answers
  - Highlighting and Note-Taking Gas Law Formula Sheet Answers
  - Interactive Elements Gas Law Formula Sheet Answers
- 8. Staying Engaged with Gas Law Formula Sheet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Gas Law Formula Sheet Answers
- 9. Balancing eBooks and Physical Books Gas Law Formula Sheet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Gas Law Formula Sheet Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Gas Law Formula Sheet Answers
  - Setting Reading Goals Gas Law Formula Sheet Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gas Law Formula Sheet Answers
  - Fact-Checking eBook Content of Gas Law Formula Sheet Answers
  - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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