1.36

1.36 Determine the mass of air in a 2 m³ tank if the air is at room temperature, 20 °C, and the absolute pressure within the tank is 200 kPa (abs).

$$m = \rho V$$
 where $V = 2m^3$ and $\rho = \rho / RT$ with $T = 20^{\circ}C = (20 + 273) K = 293 K$ and $\rho = 200 \, k \, P_4 = 200 \times 10^3 \, \frac{N}{m^2}$ Thus, $\rho = (200 \times 10^3 \, \frac{N}{m^2}) / \left[(2.869 \times 10^2 \, \frac{N \cdot m}{kg \cdot K}) (293 \, K) \right] = 2.38 \, \frac{kg}{m^3}$ Hence, $\rho = \rho V = 2.38 \, \frac{kg}{m^3} (2m^3) = \frac{4.76 \, kg}{4.76 \, kg}$

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 - o The Rise of Digital Reading Fundamentals Of Fluid Mechanics 6th Edition Solution Manual Free
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Fundamentals Of Fluid Mechanics 6th Edition Solution Manual Free
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Fundamentals Of Fluid Mechanics 6th Edition Solution Manual Free
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Fundamentals Of Fluid Mechanics 6th Edition Solution Manual Free
 - Personalized Recommendations
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 - Fundamentals Of Fluid Mechanics 6th Edition Solution Manual Free and Bestseller Lists
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 - Interactive Elements Fundamentals Of Fluid Mechanics 6th Edition Solution Manual Free

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 - Participating in Virtual Book Clubs
 - o Following Authors and Publishers Fundamentals Of Fluid Mechanics 6th Edition Solution Manual Free
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 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Fundamentals Of Fluid Mechanics 6th Edition Solution Manual Free
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Fundamentals Of Fluid Mechanics 6th Edition Solution Manual Free
 - o Setting Reading Goals Fundamentals Of Fluid Mechanics 6th Edition Solution Manual Free
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Fundamentals Of Fluid Mechanics 6th Edition Solution Manual Free
 - o Fact-Checking eBook Content of Fundamentals Of Fluid Mechanics 6th Edition Solution Manual Free
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - $\circ \ \ Utilizing \ eBooks \ for \ Skill \ Development$
 - Exploring Educational eBooks
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 - \circ Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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