

(Predicting Redox Reactions)
 Solutions

Part A:

- 1) Determine whether each of the following compounds is soluble or insoluble. For the soluble compounds, write the ions present in solution (show charges, dissociation).

- a) AgI - **insoluble**
 b) $\text{Ca}_3(\text{PO}_4)_2$ - **insoluble**
 c) CuCl_2 - **soluble**
 d) Al_2PO_4 - **soluble** $\text{AlPO}_4(\text{aq}) \rightarrow \text{Al}^{3+}(\text{aq}) + \text{PO}_4^{3-}(\text{aq})$

- 2) Complete and balance each of the following equations. If no reaction occurs, write NO REACTION.

- a) $\text{Ba}(\text{PO}_4)_2 + \text{BaCl}_2(\text{aq}) \rightarrow \text{BaCl}_2 + \text{Ba}_3(\text{PO}_4)_2$
 REACTION OF PRECIPITATION
 b) $\text{FeCl}_2(\text{aq}) + \text{KCl}(\text{aq}) \rightarrow \text{NO REACTION}$ (all reactants and possible products are soluble)
 c) $\text{FeCl}_2(\text{aq}) + \text{FeCl}_3(\text{aq}) \rightarrow \text{FeCl}_2 + \text{FeCl}_3$
 REACTION OF NO REACTION
 d) $\text{FeCl}_2(\text{aq}) + \text{FeCl}_3(\text{aq}) \rightarrow \text{FeCl}_2 + \text{FeCl}_3$
 REACTION OF NO REACTION

- 3) Write balanced complete ionic and net ionic equations for each of the following reactions:

- a) $\text{Li}_2\text{S}(\text{aq}) + \text{CuCl}_2(\text{aq}) \rightarrow \text{CuS}(\text{s}) + 2\text{LiCl}(\text{aq})$ - **balanced equation**

General Organic And Biochemistry Lab Answers

Denise Guinn



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