

3. Atoms, elements and compounds



ISOTOPE: Atom of the same element with the same proton number but a different neutron number

Have the same properties due to the same number of electrons on the outer shell

Radioisotope

- testing for leaks in pipes
- treating cancer

IONIC BONDING

Bonding: oppositely charged ions are strongly attracted by electrostatic force + involves electron transfer

metal + non metal

metals give electrons to non-metals

properties: strong bonds due to attractive forces

- conduct electricity
- high mp/bp
- charged particles
- soluble in water

macromolecular compounds

SILICON (IV) OXIDE

properties

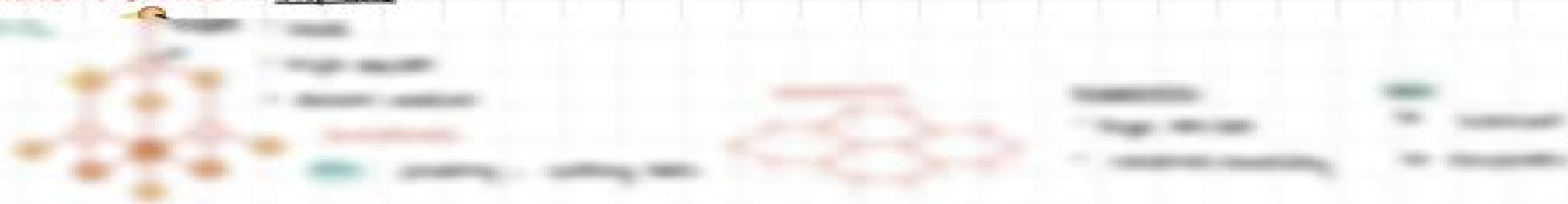
COVALENT BONDING

Bonding: Nuclei of atoms and the shared pair of electrons are strongly attracted by electrostatic forces

non metal + non metal

properties: weak bonds = low mp/bp

- insoluble
- don't conduct



COVALENT BONDING



properties of covalent compounds

- low melting point
- low boiling point
- low density
- poor conductors of electricity

properties of ionic compounds

- high melting point
- high boiling point
- high density
- good conductors of electricity

Compound	Structure	Properties
Water (H ₂ O)	Simple molecular	Low mp/bp, poor conductor
Ammonia (NH ₃)	Simple molecular	Low mp/bp, poor conductor
Silicon(IV) oxide (SiO ₂)	Macromolecular	High mp/bp, poor conductor
Sodium chloride (NaCl)	Ionic	High mp/bp, good conductor when molten

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