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Factor	Effect on rate	Diagram	Rate of reaction	Rate of reaction
Temperature	Increases		Rate of reaction increases	Rate of reaction increases
Concentration	Increases		Rate of reaction increases	Rate of reaction increases

1. Temperature: The rate of reaction increases with an increase in temperature. This is because the molecules have more kinetic energy and move faster, leading to more frequent collisions. The rate of reaction is directly proportional to the temperature.

2. Concentration: The rate of reaction increases with an increase in the concentration of the reactants. This is because there are more molecules in a given volume, leading to more frequent collisions. The rate of reaction is directly proportional to the concentration.

3. Surface Area: The rate of reaction increases with an increase in the surface area of the reactants. This is because there are more molecules exposed to the other reactant, leading to more frequent collisions. The rate of reaction is directly proportional to the surface area.

4. Catalyst: A catalyst is a substance that speeds up a chemical reaction without being consumed in the process. It works by providing an alternative reaction pathway with a lower activation energy. The rate of reaction is directly proportional to the concentration of the catalyst.

5. Pressure: The rate of reaction increases with an increase in pressure. This is because the molecules are packed more closely together, leading to more frequent collisions. The rate of reaction is directly proportional to the pressure.

6. Light: The rate of reaction increases with an increase in light intensity. This is because the light provides energy to the molecules, leading to more frequent collisions. The rate of reaction is directly proportional to the light intensity.

7. pH: The rate of reaction increases with an increase in pH. This is because the higher pH indicates a higher concentration of hydroxide ions, which act as a catalyst. The rate of reaction is directly proportional to the pH.

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