

Solutions Key

Exponents and Polynomials

ARE YOU READY? PAGE 443

1. F
2. B
3. C
4. D
5. E
6. 4^7
7. 5^2
8. $(-10)^4$
9. x^3
10. x^5
11. 9^1
12. 3^4
 $= 3 \cdot 3 \cdot 3 \cdot 3$
 $= 81$
13. -12^3
 $= -(12 \cdot 12 \cdot 12)$
 $= -144$
14. 5^3
 $= 5 \cdot 5 \cdot 5$
 $= 125$
15. 2^5
 $= 2 \cdot 2 \cdot 2 \cdot 2 \cdot 2$
 $= 32$
16. 4^3
 $= 4 \cdot 4 \cdot 4$
 $= 64$
17. $(-1)^5$
 $= (-1)(-1)(-1)(-1)(-1)$
 $= 1$
18. 0.06
19. 2,525,000
20. 15.6
21. $6 + 3p + 14 + 9p$
 $= 6 + 14 + 3p + 9p$
 $= 20 + 12p$
22. $8y - 4x + 2y + 7x - x$
 $= 8y + 2y - 4x + 7x - x$
 $= 10y + 2x$
23. $(12 + 3w - 5) + 6w - 3 - 5w$
 $= 12 - 5 - 3 + 3w + 6w - 5w$
 $= 4 + 4w$
24. $6n - 14 + 5n$
 $= 6n + 5n - 14$
 $= 11n - 14$
25. no
26. yes; $\sqrt{81} = \sqrt{9^2} = 9$
27. yes; $\sqrt{36} = \sqrt{6^2} = 6$
28. no
29. yes; $\sqrt{100} = \sqrt{10^2} = 10$
30. yes; $\sqrt{4} = \sqrt{2^2} = 2$
31. yes; $\sqrt{1} = \sqrt{1^2} = 1$
32. no

7-1 INTEGER EXPONENTS, PAGES 446–451

CHECK IT OUT! PAGES 447–448

$$1. 5^{-3} = \frac{1}{5^3} = \frac{1}{5 \cdot 5 \cdot 5} = \frac{1}{125}$$

5^{-3} m is equal to $\frac{1}{125}$ m.

$$2a. 10^{-4} = \frac{1}{10^4} = \frac{1}{10 \cdot 10 \cdot 10 \cdot 10} = \frac{1}{10,000}$$

$$b. (-2)^{-4} = \frac{1}{(-2)^4} = \frac{1}{(-2)(-2)(-2)(-2)} = \frac{1}{16}$$

$$c. (-2)^{-5} = \frac{1}{(-2)^5} = \frac{1}{(-2)(-2)(-2)(-2)(-2)} = -\frac{1}{32}$$

$$d. -2^{-5} = -\frac{1}{2^5} = -\frac{1}{2 \cdot 2 \cdot 2 \cdot 2 \cdot 2} = -\frac{1}{32}$$

$$3a. p^{-3}$$

$$= 4^{-3}$$

$$= \frac{1}{4^3}$$

$$= \frac{1}{4 \cdot 4 \cdot 4}$$

$$= \frac{1}{64}$$

$$b. 8a^{-2}b^0$$

$$= 8(-2)^{-2}(6)^0$$

$$= 8 \cdot \frac{1}{(-2)^2} \cdot 1$$

$$= 8 \cdot \frac{1}{(-2)(-2)}$$

$$= 8 \cdot \frac{1}{4}$$

$$= 2$$

$$4a. 2r^2m^{-3} = 2 \cdot r^2 \cdot m^{-3}$$

$$= 2 \cdot 1 \cdot \frac{1}{m^3}$$

$$= \frac{2}{m^3}$$

$$b. \frac{r^{-3}}{7} = r^{-3} \cdot \frac{1}{7}$$

$$= \frac{1}{7} \cdot \frac{1}{r^3}$$

$$= \frac{1}{7r^3}$$

$$c. \frac{g^4}{h^{-5}} = g^4 \cdot \frac{1}{h^{-5}}$$

$$= g^4 \cdot h^5$$

$$= g^4h^5$$

THINK AND DISCUSS, PAGE 448

1. -2; 0; 1
- 2.

Simplifying Expressions with Negative Exponents

For a negative exponent in the numerator, move the power to the denominator and change the negative exponent to a positive exponent.
possible answer: $2^{-3} = \frac{1}{2^3}$

For a negative exponent in the denominator, move the power to the numerator and change the negative exponent to a positive exponent.
possible answer: $\frac{1}{2^{-3}} = 2^3$

EXERCISES, PAGES 449–451

GUIDED PRACTICE, PAGE 449

$$1. 10^{-7} = \frac{1}{10^7} = \frac{1}{10 \cdot 10 \cdot 10 \cdot 10 \cdot 10 \cdot 10 \cdot 10}$$

$$= \frac{1}{10,000,000} \text{ m}$$

$$10^{-7} \text{ m is equal to } \frac{1}{10,000,000} \text{ m.}$$

$$2. 6^{-2} = \frac{1}{6^2} = \frac{1}{6 \cdot 6} = \frac{1}{36} \quad 3. 3^0 = 1$$

$$4. -5^{-2} = -\frac{1}{5^2} = -\frac{1}{5 \cdot 5} = -\frac{1}{25}$$

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Lauren Gardner



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discussion and conclusion the technique recrystallization was used because it is a simple method to purify a solid compound the percentage yield for benzoic acid and acetanilide acid being high indicated the initial sample had a low concentration of impurities with the melting point being in a specific

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introduction in this lab we will be using boiling points melting points and infrared spectrometry to identify unknown compounds boiling points are affect by three key factors the first being intermolecular forces inter molecular affect boiling point based on bond strength

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determination of boiling points introduction the boiling point of a compound is the temperature at which it changes from a liquid to a gas this is a physical property often used to identify substances or to check the purity of the compound it

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clean up solid and liquid spills immediately true g conclusions after conducting the boiling point determination lab it can be concluded that the unknown substance 193 is isopropyl alcohol on account of its observed boiling point of 88 c

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conclusion after completing this experiment the melting point of three different unknown compounds was obtained using the mel temp apparatus compound a had a melting point range of 114 6 115 5 c thus revealing it to be acetanilide compound b had a melting point range of 122 6 124 5 c thus

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melting point can be used as the identity and purity of an organic solid however it is not the same as the boiling point the boiling point is a characteristic unique of pure liquids the boiling point is the temperature at which the vapor pressure of a liquid is equal to the atmospheric pressure

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the purpose of the experiment is to understand what vapour pressure is and how it changes with temperature to understand the relationship between vapour pressure at a given temperature and boiling temperature at a given pressure to determine the vapour pressure of a pure liquid at various temperatures

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exploring boiling points abstract in this experiment we are given a set of straight chain alkanes and a set of compounds to compare by following the trendline of the straight chain alkanes it could be concluded that as the molar mass of the straight chain alkanes increases their boiling points increase as well

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laboratory partners samantha and mackenzie goals □ correctly determine the boiling point of an unknown liquid □ measure the refractive index of an unknown □ use the gas chromatogram to determine the mass percentage of a mixture and

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50 °C so in order to convert an experimental measurement taken in Calgary which is at higher altitude than sea level and so is at a lower pressure to that reported for sea level higher pressure one needs to add a correction factor since the boiling point at sea level is higher than that at higher altitudes

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introduction in chemistry one may encounter an unknown substance that needs to be determined by physical properties such as boiling point melting point density solubility etc in this experiment I am to figure out the boiling point of a mystery substance in order to identify it chemical responsibility this experiment requires the handling of a Bunsen burner boiling

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there are a variety of methods by which a sample's boiling point can be determined including distillation reflux and by using a Thiele tube the most straightforward method uses a Thiele tube and has the advantage of using less than 0.5 mL of material

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