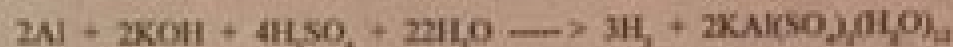


CHM 140 Lab 5 Synthesis of Alum from Aluminum Foil

Purpose: In this experiment, a commercially useful compound, alum, is synthesized from aluminum foil, KOH, and H_2SO_4 according to the following balanced reaction:



Calculating the moles of each reactant used in the synthesis will allow for the determination of the limiting reactant. Once the limiting reactant is determined, the theoretical yield of $2\text{KAl}(\text{SO}_4)_2(\text{H}_2\text{O})_{12}$ can be determined. Then, the percent yield can be calculated by dividing the actual yield by the theoretical yield and multiplying by 100.

Data from the video will be used for the calculations.

<https://youtu.be/K9LZZ1pnshI> Synthesis of Alum

Data:

1. Mass of aluminum

0.545

g

2. Mass of empty Buchner funnel and filter paper

16.153

g

3. Mass of Buchner funnel, filter paper, and alum

21.972

g

4. Mass of alum (actual yield)

5.819

g

5. Volume of 1.4 M KOH used

27.0

ml.

6. Volume of 9.0 M H_2SO_4 used

12.0

ml.

Calculations: Determine the limiting reactant. Show work for calculations.

7. Moles of Al

_____ moles Al

8. Moles of KOH

_____ moles KOH

9. Moles of H_2SO_4

_____ moles H_2SO_4

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B Lingard



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