Colorimetric Titration Experiment - Determine the amount of vitamin C in a medium peach

A suitable method for the determination of vitamin C (ascorbic acid, $C_6H_8O_6$) is a titration with potassium lodate (KIO₃). In this reaction, potassium lodate is used as a titrant and when added to an ascorbic acid solution that contains strong acid and potassium lodide (KI), the potassium lodate reacts with potassium lodide, liberating molecular lodine (I_2) as shown in the reaction below:

During the titration, as long as the solution contains ascorbic acid, the I_2 produced in equation I is used up in a rapid reaction with ascorbic acid (equation 2), during which dehydroascorbic acid ($C_4H_4O_4$) and iodide ion (I) are formed:

[2]
$$C_6H_6O_6 + I_2 \rightarrow C_6H_6O_6 + 2\Gamma + 2H^*$$

Once all the ascorbic acid has been consumed, any excess iodine (I₂) will remain in solution. This excess iodine reacts with starch, to form an intensely blue colored complex, indicating that the endpoint is reached.

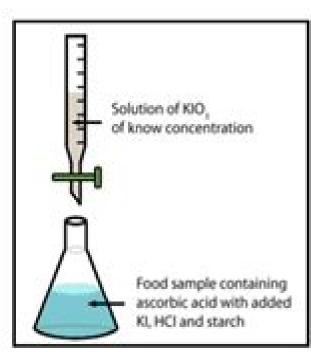
The experimental setup is shown schematically to the right.

The virtual lab stockroom contains the following solutions:

Solid potassium lodate - (KIO₃). You will need to create a stock solution with this in order to perform your titration. (Note: The molar mass of potassium lodate is 214.001 g/mol.)

Potassium lodide solution - This is provided as a concentrated solution of KI in water. Since you are required to use an excess KI, the amount does not need to be measured precisely. Use 5 mL of KI solution for each titration.

2M HCI - The addition of acid is necessary to provide the acidic conditions required in the iodometric titration (see Equation 1). Use 4mL of the HCI solution in each titration.



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D Keegan

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Nutrition Labeling Handbook Ralph Shapiro,1995-08-30 This handbook examines the Nutritional Labeling and Education Act NLEA passed by Congress in 1990 It discusses the history of the NLEA and its impact on various segments of the food industry making complex and detailed regulations easily understandable throughout Government industry and consumer perspectives on labelling regulations are provided along with practical guidelines for compliance and packaging

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Illustrated Guide to Home Chemistry Experiments Robert Bruce Thompson, 2012-02-17 For students DIY hobbyists and science buffs who can no longer get real chemistry sets this one of a kind guide explains how to set up and use a home chemistry lab with step by step instructions for conducting experiments in basic chemistry not just to make pretty colors and stinky smells but to learn how to do real lab work Purify alcohol by distillation Produce hydrogen and oxygen gas by electrolysis Smelt metallic copper from copper ore you make yourself Analyze the makeup of seawater bone and other common substances Synthesize oil of wintergreen from aspirin and rayon fiber from paper Perform forensics tests for fingerprints blood drugs and poisons and much more From the 1930s through the 1970s chemistry sets were among the most popular Christmas gifts selling in the millions But two decades ago real chemistry sets began to disappear as manufacturers and retailers became concerned about liability em The Illustrated Guide to Home Chemistry Experiments steps up to the plate with lessons on how to equip your home chemistry lab master laboratory skills and work safely in your lab The bulk of this book consists of 17 hands on chapters that include multiple laboratory sessions on the following topics Separating Mixtures Solubility and Solutions Colligative Properties of Solutions Introduction to Chemical Reactions Stoichiometry Reduction Oxidation Redox Reactions Acid Base Chemistry Chemical Kinetics Chemical Equilibrium and Le Chatelier s Principle Gas Chemistry Thermochemistry and Calorimetry Electrochemistry Photochemistry Colloids and Suspensions Qualitative Analysis Quantitative Analysis Synthesis of Useful Compounds Forensic Chemistry With plenty of full color illustrations and photos Illustrated Guide to Home Chemistry Experiments offers introductory level sessions suitable for a middle school or first year high school chemistry laboratory course and more advanced sessions suitable for students who intend to take the College Board Advanced Placement AP Chemistry exam A student who completes all of the laboratories in this book will have done the equivalent of two full years of high school chemistry lab work or a first year college general chemistry laboratory course This hands on introduction to real chemistry using real equipment real chemicals and real quantitative experiments is ideal for the many thousands of young people and adults who want to experience the magic of chemistry **Photoacoustic IR Spectroscopy** Kirk H. Michaelian, 2010-12-01 This invaluable and up to date source on instruments and applications covers everything needed to employ a technique for investigating various gases and materials including biomaterials It includes the latest developments in light sources signal recovery and numerical methods There is no other single publication that reviews the entire subject of photoacoustic infrared spectroscopy in such detail Physicists chemists and spectroscopists in both academic and industrial laboratories polymer and organic chemists analysts in industry forensic and government laboratories and materials scientists will find this book to be a vital resource The Analytical Chemistry Laboratory Companion Michael D. Holloway, 2025-07-15 The Analytical Chemistry Laboratory Companion is

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