Lab - Evidence for Chemical Change

One way of knowing that a chemical change has occurred is to observe that the properties of the products are different from those of the reactants. The new product can then become a reactant in another chemical reaction. In this experiment you will observe a sequence of changes that occur when a solution that begins as copper (II) nitrate is treated with a series of different reactants. All of the reactions will take place in the same test tube. At each step you will look for evidence that a new substance is formed as a result of a chemical change. You will also observe heat changes and relate them to chemical reactions.

Pre-Lab

For each reaction: **IDENTIFY** the type of chemical reaction, **PREDICT** the products, and **BALANCE** the chemical equation.

1. _ Cu(NO₃)₂ (aq) + _ NaOH (aq) \rightarrow

Type of Reaction:

Cu(OH)₂ (s) + heat →

Type of Reaction:

CuO (s) + _ HCl (aq) →

Type of Reaction:

CuCl₂ (aq) + __ Al (s) →

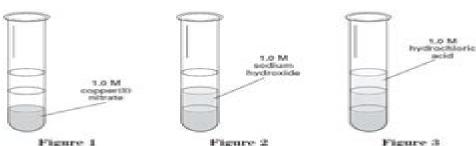
Type of Reaction:

HCl (aq) + __ Al (s) →

Type of Reaction:

Procedure

- Place 50 mL of water into the 100 mL beaker and heat it on a hot plate until boiling. This will be the hot-water bath you will use in step 6.
- While the water bath is heating, use the lab marker and ruler to make three marks on the test tube that are 1 cm apart. Make the marks starting at the bottom of the test tube and moving toward the top.
- Add 1.0 M Cu(NO₃)₂, copper (II) nitrate, to the **first** mark on the test tube as in **Figure** 1.
- Add 1.0 M NaOH, sodium hydroxide, up to the second mark on the test tube as in Figure 2.



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Guillaume Favre

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