

Double Replacement Reactions and Solubility

Post-Lab Questions: Review your Observations Table. Complete AND balance each equation below. If your results yielded a precipitate, then write the net ionic equation in the box to the right. MAKE SURE TO SHOW STATES/PHASES for every reactant and product! Remember that all nitrates are soluble in water. If you did not see signs of a precipitate, then write NONE in the box to the right.

1) $\text{Na}_2\text{CO}_{3(aq)} + \text{Al}(\text{NO}_3)_3(aq) \rightarrow$	$6\text{NaNO}_3(aq) + \text{Al}(\text{OH})_3(s)$	None
2) $\text{NaCl}_{(aq)} + \text{NH}_4\text{NO}_3(aq) \rightarrow$	$\text{NaNO}_3(aq) + \text{NH}_4\text{Cl}(aq)$	None
3) $\text{NaI}_{(aq)} + \text{AgNO}_{3(aq)} \rightarrow$	$\text{NaNO}_3(aq) + \text{AgI}(s)$	$\text{Ag}^+(aq) + \text{I}^-(aq) \rightarrow \text{AgI}(s)$
4) $\text{Na}_3\text{PO}_{4(aq)} + \text{Cu}(\text{NO}_3)_2(aq) \rightarrow$	$6\text{NaNO}_3(aq) + (\text{Cu}_3(\text{PO}_4)_2)(s)$	$3\text{Cu}^{2+}(aq) + 2\text{PO}_4^{3-}(aq) \rightarrow \text{Cu}_3(\text{PO}_4)_2(s)$

Use your observations table to answer the following. *ALL ions must include a symbol AND a charge or it is wrong!*

5 a) The symbols and charges of the anions that produced the fewest (<3) precipitates are _____

b) Salts of these anions may be described as SOLUBLE except for those with the cations _____

6a) The symbols and charges of the anion that produced the most (>5) precipitates are _____

b) Salts of these anions may be described as INSOLUBLE except for those with the cations _____

7. Although you did not test the bromide ion (Br^-), would you expect most bromide salts (AlBr_3 , BaBr_2 , CaBr_2 , etc) to be soluble or insoluble? Hint: Locate bromine on the periodic table. _____

8. Explain your answer in #7 _____

9. What cation combined with bromide is an exception to the rule in #7? _____

10-14 Summarize the general solubility rules of ionic compounds in the table below. Two have been done for you.

Ion	Generally soluble or insoluble?	Exceptions
9) carbonates		
10) chlorides, bromides, iodides		
11) hydroxides		
nitrates	Soluble	None
12) phosphates		

Double Replacement Reaction Lab Answers

Kristin Tuttle Bump



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Chemistry in the Laboratory James M. Postma, Julian L. Roberts, J. Leland Hollenberg, 2009-08-07 For nearly 40 years Chemistry in the Laboratory has been meeting the needs of teachers and students This new edition builds on that legacy while addressing cutting edge trends in the chemistry laboratory including forensic chemistry and environmental and green chemistry As always the new edition of Chemistry in the Laboratory offers precise easy to follow instructions helpful illustrations and an emphasis throughout on laboratory safety Again throughout a Consider This feature encourages users to expand the principles of the experiment into interesting applications open ended experiments or unexplored corners Most experiments in the manual can be completed in one lab session but some can be linked or extended for a multi lab project

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Table of Contents Double Replacement Reaction Lab Answers

1. Understanding the eBook Double Replacement Reaction Lab Answers
 - The Rise of Digital Reading Double Replacement Reaction Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Double Replacement Reaction Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Double Replacement Reaction Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Double Replacement Reaction Lab Answers
 - Personalized Recommendations
 - Double Replacement Reaction Lab Answers User Reviews and Ratings
 - Double Replacement Reaction Lab Answers and Bestseller Lists
5. Accessing Double Replacement Reaction Lab Answers Free and Paid eBooks
 - Double Replacement Reaction Lab Answers Public Domain eBooks
 - Double Replacement Reaction Lab Answers eBook Subscription Services

- Double Replacement Reaction Lab Answers Budget-Friendly Options
- 6. Navigating Double Replacement Reaction Lab Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Double Replacement Reaction Lab Answers Compatibility with Devices
 - Double Replacement Reaction Lab Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Double Replacement Reaction Lab Answers
 - Highlighting and Note-Taking Double Replacement Reaction Lab Answers
 - Interactive Elements Double Replacement Reaction Lab Answers
- 8. Staying Engaged with Double Replacement Reaction Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Double Replacement Reaction Lab Answers
- 9. Balancing eBooks and Physical Books Double Replacement Reaction Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Double Replacement Reaction Lab Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Double Replacement Reaction Lab Answers
 - Setting Reading Goals Double Replacement Reaction Lab Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Double Replacement Reaction Lab Answers
 - Fact-Checking eBook Content of Double Replacement Reaction Lab Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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