

Name: _____

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Experiment 8.6: Status of Iron and Copper

Book/Book Chemistry Laboratory Experiments, pp. 124-125

The student's experiment will test separating chemical reactions. The main suggested test is a 2.0M solution, to separate a mixture of particles. Other important features: the main idea of any separation or solution is to change its physical characteristics (the chemical formula, or chemical mass, being used). The main reason to test is, which means, that the main idea, mass is being applied to it.

The student, for experimental purposes, chemical reactions, for this experiment, you will observe the reaction of iron metal with a solution of copper(II) sulfate and observe the reaction of iron metal with a solution of iron(II) sulfate.

Objectives:

- to determine the number of moles of copper produced in the reaction of iron metal with a solution of copper(II) sulfate
- to determine the number of moles of iron produced in the reaction of iron metal with a solution of copper(II) sulfate
- to determine the number of moles of iron produced in the reaction of iron metal with a solution of iron(II) sulfate
- to determine the number of moles of iron produced in the reaction of iron metal with a solution of iron(II) sulfate

Procedure:

In this experiment, you will observe the reaction of iron metal with a solution of copper(II) sulfate. The reaction is described by the chemical equation in the solution and the reaction of iron metal with a solution of iron(II) sulfate. The reaction is described by the chemical equation in the solution and the reaction of iron metal with a solution of iron(II) sulfate. The reaction is described by the chemical equation in the solution and the reaction of iron metal with a solution of iron(II) sulfate.

After the reaction, the main idea is to separate the reaction and the reaction. The reaction is described by the chemical equation in the solution and the reaction of iron metal with a solution of iron(II) sulfate. The reaction is described by the chemical equation in the solution and the reaction of iron metal with a solution of iron(II) sulfate.

In order to find the number of moles of iron produced, you will observe the reaction of iron metal with a solution of iron(II) sulfate. The reaction is described by the chemical equation in the solution and the reaction of iron metal with a solution of iron(II) sulfate.

Data and Calculations:

Before the reaction:	Mass of copper (g): _____
	Mass of copper(II) sulfate: _____
	Mass of iron (g): _____
After the reaction:	Mass of iron (g): _____
	Mass of iron(II) sulfate (g): _____

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Padhraic Smyth



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