

Chem 331, Exam 1

Friday, September 21, 2007 (100 points total)

$$R = 8.3145 \text{ J K}^{-1} \text{ mol}^{-1} = 0.082058 \text{ L atm K}^{-1} \text{ mol}^{-1}$$

$$1 \text{ atm} = 101.325 \text{ kPa} = 760 \text{ Torr} = 1.01325 \text{ bar} \quad (1 \text{ bar} = 10^5 \text{ Pa}); \quad 1 \text{ L} = 1 \text{ dm}^3$$

$$\mu_{J-T} = \left(\frac{\partial T}{\partial P} \right)_H, \quad \left(\frac{T_f}{T_i} \right) = \left(\frac{V_f}{V_i} \right)^{\gamma - \frac{C_p}{C_v}}, \quad w = -nRT \ln \frac{V_f}{V_i}$$

Problems (20 pts each) Show all work

- 1.) Two moles of an ideal monatomic gas initially at a pressure of 1.00 bar and 25°C is compressed adiabatically and irreversibly by application of a constant external pressure of 15.00 bar, until the final pressure of the gas reaches 10.00 bar. Assume the gas has molar heat capacity $\bar{C}_v = 3R/2$ and is independent of temperature.

- (a) Calculate the heat absorbed by the gas.
 (b) Calculate the final temperature of the gas.

a) $q = 0$ (adiabatic) T_f
 b) $\Delta U = w \Rightarrow \int_{T_i}^{T_f} n \bar{C}_v dT = \int_{V_i}^{V_f} -P_{\text{ext}} dV$ (since ideal gas)

$$n \bar{C}_v (T_f - T_i) = -P_{\text{ext}} (V_f - V_i)$$

$$= -P_{\text{ext}} \left[\frac{nRT_f}{P_f} - \frac{nRT_i}{P_i} \right]$$

rearrange:

$$T_f = T_i \frac{n \bar{C}_v + nR \frac{P_{\text{ext}}}{P_i}}{n \bar{C}_v + nR \frac{P_{\text{ext}}}{P_f}}$$

$$= 298 \frac{(2 \cdot 3R/2) + (2)(R)(15)}{(2 \cdot 3R/2) + (2)(R)(10)}$$

$$= 1639 \text{ K}$$

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Jin-Ying Zhang



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