Chapter 15. FLUIDS

Density:

15.1. What volume does 0.4 kg of alcohol occupy? What is the weight of this volume?

$$\rho = \frac{m}{V}$$
; $V = \frac{m}{\rho} = \frac{0.4 \text{ kg}}{790 \text{ kg/m}^2}$; $V = 5.06 \times 10^{-4} \text{ m}^3$

 $W = DV = \rho_S V = 790 \text{ kg/m}^3)(9.8 \text{ m/s}^2)(5.06 \times 10^{-4} \text{ m}^3); W = 3.92 \text{ N}$

15-2. An unknown substance has a volume of 20 ft³ and weighs 3370 lb. What are the weight density and the mass density?

$$D = \frac{W}{V} = \frac{3370 \text{ lb}}{20 \text{ ft}^3}; \quad D = 168 \text{ lb/ft}^3 \qquad \rho = \frac{D}{g} = \frac{168 \text{ lb/ft}^3}{9.8 \text{ m/s}^3}; \quad \rho = 5.27 \text{ slugs/ft}^3$$

15-3. What volume of water has the same mass of 100 cm³ of lead? What is the weight density of lead?

First find mass of lead: $m = \rho V = (11.3 \text{ g/cm}^3)(100 \text{ cm}^3)$; m = 1130 g

Now water:
$$V_w = \frac{m}{\rho_w} = \frac{1130 \text{ g}}{1 \text{ g/cm}^2} = 1130 \text{ cm}^2$$
: $V_w = 1130 \text{ cm}^3$ $D = \rho_S$

 $D = (11,300 \text{ kg/m}^3)(9.8 \text{ m/s}^2) = 110,740 \text{ N/m}^3;$ $D = 1.11 \times 10^3 \text{ N/m}^3$

15-4. A 200-mL flask (1 L = 1 x 10⁻³ m³) is filled with an unknown liquid. An electronic balance indicates that the added liquid has a mass of 176 g. What is the specific gravity of the liquid? Can you guess the identity of the liquid?

$$V = 200 \text{ mL} = 0.200 \text{ L} = 2 \text{ x } 10^{-4} \text{ m}^3$$
; $m = 0.176 \text{ kg}$

$$\rho = \frac{m}{V} = \frac{0.176 \text{ kg}}{2 \times 10^{-4} \text{m}^3}$$
; $\rho = 880 \text{ kg/m}^3$, Benzene

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Bruce R. Munson, Donald F.
Young, Theodore H. Okiishi, Wade W.
Huebsch

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