

Determination of the Solubility-Product Constant for a Sparingly Soluble Salt

A. Preparation of a Calibration Curve

Initial $[\text{CrO}_4^{2-}]$ 0.0024

	Volume of 0.0024 M K_2CrO_4	Total volume	$[\text{CrO}_4^{2-}]$	Absorbance
1.	<u>1 mL</u>	<u>100 mL</u>	<u> </u>	<u>0.122</u>
2.	<u>5 mL</u>	<u>100 mL</u>	<u> </u>	<u>0.625</u>
3.	<u>10 mL</u>	<u>100 mL</u>	<u> </u>	<u>1.288</u>
4.	<u>15 mL</u>	<u>100 mL</u>	<u> </u>	<u>1.955</u>

Molar absorption coefficient for $[\text{CrO}_4^{2-}]$ Error for molar absorption coefficient

B. Determination of the Solubility-Product Constant

	Absorbance	$[\text{CrO}_4^{2-}]$	$[\text{Ag}^+]$	K_{sp} of $\text{Ag}_2\text{CrO}_4 = [\text{CrO}_4^{2-}][\text{Ag}^+]^2$
1.	<u>1.357</u>	<u> </u>	<u> </u>	<u> </u>
2.	<u>1.327</u>	<u> </u>	<u> </u>	<u> </u>
3.	<u>1.549</u>	<u> </u>	<u> </u>	<u> </u>

Determination Of A Solubility Product Lab Answers

William Shurtleff; Akiko Aoyagi



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