

KEY

1. Complete and classify the following partial equations

- a) $1 \text{ Fe} + 3 \text{H}_2\text{SO}_4 \rightarrow 1 \text{ FeSO}_4 + 3 \text{H}_2$ (single repl)
 b) $\text{NaI} + \text{Cl}_2 \rightarrow$
 c) $1 \text{ Zn} + 1 \text{ Pb(NO}_3)_2 \rightarrow 1 \text{ Zn(NO}_3)_2 + 1 \text{ Pb}$ (single repl)
 d) $\text{Al}_2\text{O}_3 + \text{HCl} \rightarrow$
 e) $2 \text{ Cu} + \text{O}_2 \rightarrow 2 \text{ CuO}$ or $4 \text{ Cu} + 1 \text{ O}_2 \rightarrow 2 \text{ Cu}_2\text{O}$ (synthesis)
 f) $\text{Al}_2\text{O}_3 \rightarrow$
 g) $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$ (synth)
 h) $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow$
 i) $1 \text{ Mg} + 2 \text{ HNO}_3 \rightarrow 1 \text{ Mg(NO}_3)_2 + 1 \text{ H}_2$ (s.r.)
 j) $\text{Na} + \text{H}_2\text{O} \rightarrow$
 k) $3 \text{ NaOH} + \text{H}_3\text{PO}_4 \rightarrow 1 \text{ Na}_3\text{PO}_4 + 3 \text{ H}_2\text{O}$ (acid base)
 l) $\text{H}_2\text{SO}_4 + \text{Al(OH)}_3 \rightarrow$
 m) $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow 2 \text{ CO}_2 + 3 \text{ H}_2\text{O}$ (combustion)
 n) $\text{AlCl}_3 \rightarrow$
 o) $2 \text{ KOH} + \text{H}_2\text{SO}_4 \rightarrow 1 \text{ K}_2\text{SO}_4 + 2 \text{ H}_2\text{O}$ (acid base)
 p) $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow$
 q) $1 \text{ NaOH} + \text{HCl} \rightarrow 1 \text{ NaCl} + 1 \text{ H}_2\text{O}$ (acid base)
 r) $\text{NaHCO}_3 + \text{H}_2\text{SO}_4 \rightarrow$
 s) $1 \text{ C}_2\text{H}_6 + 5 \text{ O}_2 \rightarrow 2 \text{ CO}_2 + 3 \text{ H}_2\text{O}$ (combustion)
 t) $\text{H}_2\text{O} \rightarrow$
 u) $2 \text{ HNO}_3 + \text{CaCO}_3 \rightarrow 1 \text{ Ca(NO}_3)_2 + 1 \text{ CO}_2 + 1 \text{ H}_2\text{O}$ (s.r.)
 v) $\text{CH}_4 + \text{O}_2 \rightarrow$
 w) $1 \text{ Pb} + 1 \text{ H}_2\text{SO}_4 \rightarrow 1 \text{ PbSO}_4 + 1 \text{ H}_2$ (s.r.)
 x) $\text{KI} + \text{Br}_2 \rightarrow$
 y) $1 \text{ Ba(OH)}_2 + 2 \text{ HNO}_3 \rightarrow 1 \text{ Ba(NO}_3)_2 + 2 \text{ H}_2\text{O}$ (acid base)
 z) $\text{Pb} + \text{Cu(NO}_3)_2 \rightarrow$

2. Given: $2 \text{ NaClO}_3 \rightarrow 2 \text{ NaCl} + 3 \text{ O}_2$ calculate the mass of oxygen gas produced if 143 g of NaClO_3 is decomposed.

65.4g

3. Calculate the mass of oxygen gas required to burn 242 g of wax ($\text{C}_{21}\text{H}_{44}$)

254g

4. If 14.0 moles of butane (C_4H_{10}) is burned

- a) calculate the number of moles of oxygen gas required. 90.0 mol
 b) calculate the number of moles of carbon dioxide produced. 56.0 mol
 c) calculate the mass of water produced. 1080 g

5. Metallic aluminum reacts with zinc nitrate to produce metallic zinc and aluminum nitrate.

- a) Calculate the mass of aluminum that will react with 1.30 moles of zinc nitrate. 27.4g Al
 b) Calculate the mass of zinc that would be produced from 50.0 g of aluminum and sufficient zinc nitrate. 182 g Zn

6. Given: $\text{Cu} + 2 \text{ AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2 \text{ Ag}$

- a) How many grams of Cu are required to react with 3.00 g of silver nitrate? 0.544g Cu
 b) How many grams of silver are produced from 3.00 g of silver nitrate and sufficient Cu? 1.70g Ag

2. $2 \text{ NaClO}_3 \rightarrow 2 \text{ NaCl} + 3 \text{ O}_2$
 143g NaClO_3
 143g $\times \frac{1 \text{ mol NaClO}_3}{122.44 \text{ g}} = 1.17 \text{ mol NaClO}_3$
 $1.17 \text{ mol NaClO}_3 \times \frac{3 \text{ mol O}_2}{2 \text{ mol NaClO}_3} = 1.75 \text{ mol O}_2$
 $1.75 \text{ mol O}_2 \times 32 \text{ g/mol} = 56 \text{ g O}_2$

3. $\text{C}_{21}\text{H}_{44} + 32 \text{ O}_2 \rightarrow 21 \text{ CO}_2 + 22 \text{ H}_2\text{O}$
 242g $\text{C}_{21}\text{H}_{44}$
 $242 \text{ g} \times \frac{1 \text{ mol}}{284 \text{ g}} = 0.852 \text{ mol}$
 $0.852 \text{ mol} \times 32 \text{ mol O}_2 = 27.26 \text{ mol O}_2$
 $27.26 \text{ mol O}_2 \times 32 \text{ g/mol} = 872 \text{ g O}_2$

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JL Elias

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 - The Rise of Digital Reading Nelson Chemistry 30 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Nelson Chemistry 30 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nelson Chemistry 30 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Nelson Chemistry 30 Answer Key
 - Personalized Recommendations
 - Nelson Chemistry 30 Answer Key User Reviews and Ratings
 - Nelson Chemistry 30 Answer Key and Bestseller Lists
5. Accessing Nelson Chemistry 30 Answer Key Free and Paid eBooks
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 - Nelson Chemistry 30 Answer Key eBook Subscription Services
 - Nelson Chemistry 30 Answer Key Budget-Friendly Options

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 - Nelson Chemistry 30 Answer Key Enhanced eBook Features
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 - Highlighting and Note-Taking Nelson Chemistry 30 Answer Key
 - Interactive Elements Nelson Chemistry 30 Answer Key
8. Staying Engaged with Nelson Chemistry 30 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Nelson Chemistry 30 Answer Key
9. Balancing eBooks and Physical Books Nelson Chemistry 30 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Nelson Chemistry 30 Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Nelson Chemistry 30 Answer Key
 - Setting Reading Goals Nelson Chemistry 30 Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Nelson Chemistry 30 Answer Key
 - Fact-Checking eBook Content of Nelson Chemistry 30 Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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