

11-3 Review and Reinforcement (continued)

10. Hydrogen is produced when methane (CH_4) reacts with water. The other product is CO_2 . Using 80.0 g of methane and 16.3 g of water, how many liters of H_2 can be produced at STP? What is the limiting reactant?
11. Methyl alcohol (CH_3OH) is made by reacting carbon monoxide with H_2 . Starting with 2.5 g of H_2 and 30.0 L of CO , what mass of methyl alcohol could be produced at STP? Which is the limiting reactant?
12. In a reaction of 15.3 g of NaCl with 60.8 g of $\text{Pb}(\text{NO}_3)_2$, how many grams of lead(II) chloride will be produced? What is the limiting reactant?
13. Diborane (B_2H_6) is widely used in the synthesis of organic compounds. Diborane itself is made by the reaction $2\text{NaBH}_4 + \text{I}_2 \rightarrow \text{B}_2\text{H}_6 + 2\text{NaI} + \text{H}_2$. If 6.3 g of NaBH_4 are reacted with excess I_2 , how many grams of diborane could theoretically be isolated? If 1.9 g of diborane is actually produced, what is the percent yield of the reaction?
14. Ammonia gas can be produced by reacting CaO with NH_4Cl . Water and calcium chloride are produced in addition to ammonia. If 23.0 g of CaO and 50.0 g of NH_4Cl are mixed, what is the maximum possible volume of NH_3 that can be produced at STP? If 16.1 L of NH_3 are actually produced, what is the percent yield of NH_3 ?

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N Colangelo



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