

**10.213 Chemical Engineering Thermodynamics
Spring 2002**

Problem Set F

Due Friday, March 22, 2002 in lecture

Reading Assignments from SVN&A:

Mon 3/11:	6.2-6.6
Wed 3/13:	6.2, 6.7
Fri 3/15:	2.12, pg 257-273
Mon 3/18:	8-8.1
Wed 3/20:	
Fri 3/22:	8.2
Mon 4/1:	9-9.5

Note: In reporting thermodynamic quantities obtained from graphs, charts, and/or tables on your problem sets and tests, please note the relevant conditions. As an example: a solution might note "steam at $T_1 = 300\text{ }^\circ\text{C}$ and $P_1 = 100\text{ kPa}$, $H_1 = 3074.5\text{ kJ/kg}$ (steam table)." Incorrect values that cannot be tracked down will receive no partial credit.

Problem 20

Five moles of calcium carbide are combined with 10 moles of liquid water in a closed, rigid, high-pressure vessel of 750 mL capacity. Acetylene gas is produced by the reaction:



Initial conditions are $25\text{ }^\circ\text{C}$ and 1 bar, and the reaction goes to completion. For a final temperature of $125\text{ }^\circ\text{C}$, determine the final pressure in the chamber and the heat transferred. At $125\text{ }^\circ\text{C}$, the molar volume of $\text{Ca}(\text{OH})_2$ is $33.0\text{ cm}^3/\text{mol}$. Ignore the effect of any gas present in the tank initially.

Problem 21

- A rigid vessel of 2-m^3 capacity contains 0.02 m^3 of liquid water and 1.98 m^3 of water vapor at 101.33 kPa . How much heat must be added to the contents of the vessel so that the liquid water is just evaporated?
- A turbine operates adiabatically with superheated steam entering at 45 bar and $400\text{ }^\circ\text{C}$. If the exhaust steam must be "dry", what is the minimum allowable exhaust pressure for a turbine efficiency of 0.75 ? What would be the work output for the turbine? Suppose the efficiency was 0.80 . Would the minimum exhaust pressure be higher or lower? Explain.

Problem 22

Ammonia is compressed from $21\text{ }^\circ\text{C}$ and 200 kPa to 1 MPa in an adiabatic compressor with an efficiency of 0.82 . Estimate the final temperature, the work required, and the entropy change of the ammonia

- By generalized correlations
- By use of <http://webbook.nist.gov/chemistry/fluid/>

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Lei Shi



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Refrigerating Engineering ,1925 Vols 1 17 include Proceedings of the 10th 24th 1914 28 annual meeting of the society

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web the mole is an amount unit similar to familiar units like pair dozen gross etc it provides a specific measure of the number of atoms or molecules in a sample of matter one latin connotation for the word mole is large mass or bulk which is consistent with its use as the name for this unit

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web below are some important points to get clear mole concept definition of mole chemistry and mole units mole can be define as the atomic mass molecular mass ionic mass or formula mass of a substance expressed in grams is called mole of that substance

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web solution summary learning objectives describe the unit mole relate the mole quantity of substance to its mass so far we have been talking about chemical substances in terms of individual atoms and molecules

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web libretexts the mole concept is central to any chemistry calculation based on experimental results the mole is how we relate the unbelievably small atoms and molecules that make something up to the measurable properties such as mass which we may observe in a laboratory setting

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web nov 23 2022 mole concept is the method used to express the amount of substance this has been experimentally proving that one gram atom of any element as well as one gram molecule of any substance contains the same amount of entities the experimentally decided number is found to be 6.022137×10^{23}

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web oct 10 2023 mole standard unit $6.02214076 \times 10^{23}$ in chemistry for measuring large quantities of very small entities such as atoms molecules or other specified particles the number of units in a mole also bears the name avogadro s number or avogadro s constant in honor of the italian physicist amedeo avogadro

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web the mole was made the seventh si base unit in 1971 by the 14th cgm 2019 redefinition of si base units before the 2019 redefinition of the si base units the mole was defined as the amount of substance of a system that contains as many elementary entities as there are atoms in 12 grams of carbon 12 the most common isotope of carbon