

Name _____

5-3 Review and Reinforcement

Periodic Trends

Use the periodic table and your knowledge of periodic trends to answer the following questions.

Which atom in each pair has the larger atomic radius?

- | | |
|-----------|-------------|
| <u>K</u> | 1. Li or K |
| <u>Ca</u> | 2. Ca or Na |
| <u>Ge</u> | 3. Ge or B |
| <u>O</u> | 4. O or C |
| <u>Br</u> | 5. Cl or Br |
| <u>Ba</u> | 6. Be or Ba |
| <u>Si</u> | 7. Si or S |
| <u>Au</u> | 8. Fe or Au |

Which ion in each pair has the smaller atomic radius?

- | | |
|-----------------------------|----------------------------|
| <u>K^+</u> | 9. K^+ or Cl^- |
| <u>Be^{2+}</u> | 10. Be^{2+} or F^- |
| <u>Al^{3+}</u> | 11. Al^{3+} or P^{3-} |
| <u>K^+</u> | 12. K^+ or Ca^2+ |
| <u>Fe^{3+}</u> | 13. Fe^{2+} or Fe^{3+} |
| <u>F^-</u> | 14. F^- or Cl^- |

Which atom or ion in each pair has the largest ionization energy?

- | | |
|-----------|--------------|
| <u>O</u> | 15. Na or O |
| <u>Be</u> | 16. Be or Ba |
| <u>F</u> | 17. Ar or F |
| <u>Cu</u> | 18. Cu or Fe |
| <u>Ne</u> | 19. I or Ne |
| <u>V</u> | 20. K or V |
| <u>Ca</u> | 21. Ca or Fr |
| <u>Se</u> | 22. W or Se |

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Mike Jess



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Table of Contents 11 3 Review And Reinforcement Answer Key

1. Understanding the eBook 11 3 Review And Reinforcement Answer Key
 - The Rise of Digital Reading 11 3 Review And Reinforcement Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying 11 3 Review And Reinforcement Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 11 3 Review And Reinforcement Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from 11 3 Review And Reinforcement Answer Key
 - Personalized Recommendations
 - 11 3 Review And Reinforcement Answer Key User Reviews and Ratings
 - 11 3 Review And Reinforcement Answer Key and Bestseller Lists
5. Accessing 11 3 Review And Reinforcement Answer Key Free and Paid eBooks
 - 11 3 Review And Reinforcement Answer Key Public Domain eBooks
 - 11 3 Review And Reinforcement Answer Key eBook Subscription Services
 - 11 3 Review And Reinforcement Answer Key Budget-Friendly Options

6. Navigating 11 3 Review And Reinforcement Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - 11 3 Review And Reinforcement Answer Key Compatibility with Devices
 - 11 3 Review And Reinforcement Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 11 3 Review And Reinforcement Answer Key
 - Highlighting and Note-Taking 11 3 Review And Reinforcement Answer Key
 - Interactive Elements 11 3 Review And Reinforcement Answer Key
8. Staying Engaged with 11 3 Review And Reinforcement Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 11 3 Review And Reinforcement Answer Key
9. Balancing eBooks and Physical Books 11 3 Review And Reinforcement Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 11 3 Review And Reinforcement Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 11 3 Review And Reinforcement Answer Key
 - Setting Reading Goals 11 3 Review And Reinforcement Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 11 3 Review And Reinforcement Answer Key
 - Fact-Checking eBook Content of 11 3 Review And Reinforcement Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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