

Stoichiometry

Quiz 11*9 Section 9

Section 9: Introduction to Stoichiometry

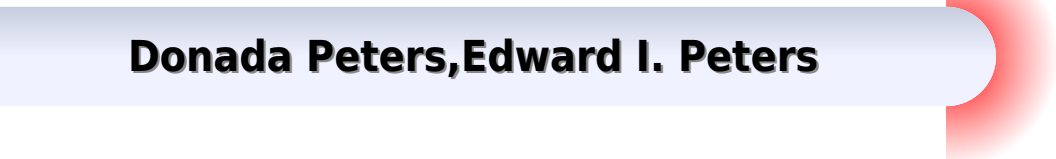
In the space provided, write the letter of the best response that best completes each problem or best answers each question.

1. Stoichiometry is the branch of chemistry that deals with, according to convention, and with molecular relationships, chemical reactions,
 - a. balancing
 - b. balancing
 - c. energy conversion
 - d. mass relationships
 - e. physical measurements
2. Some three-body reactions are favored at elevated temperatures,
 - a. the forward reaction is endothermic.
 - b. the total gas loss.
 - c. the rate of concentration is not temperature.
 - d. temperature decreases.
3. The number of significant figures in the product of two numbers is determined by
 - a. the number of significant figures in the factor with the least number of significant figures.
 - b. the number of decimal places in the factor having the most decimal places.
 - c. the number of significant figures of the least significant number in the product.
 - d. the number of significant figures in the number having the most significant figures.
4. The chemical equation $\text{P}_4\text{O}_{10} + 6\text{H}_2\text{O} \rightarrow 4\text{H}_3\text{PO}_4$ can be interpreted correctly as
 - a. 1 molecule of P_4O_{10} + 6 molecules of H_2O forms 4 molecules of H_3PO_4 .
 - b. 1 mole of P_4O_{10} + 6 moles of H_2O forms 4 moles of H_3PO_4 .
 - c. 1 g of P_4O_{10} + 6 g of H_2O forms 4 g of H_3PO_4 .
 - d. both a and b.
5. How many moles of H_2O are produced from the following chemical equation?
 - a. 1
 - b. 2
 - c. 3
 - d. 4



Section Quiz Introduction To Stoichiometry Answers

Donada Peters, Edward I. Peters



Section Quiz Introduction To Stoichiometry Answers:

Chemtrek Stephen Thompson,1990 *Introduction to Chemical Processes: Principles, Analysis, Synthesis* Regina M. Murphy,2007 Introduction to Chemical Processes Principles Analysis Synthesis enhances student understanding of the connection between the chemistry and the process Users will find strong coverage of chemistry gain a solid understanding of what chemical processes do convert raw materials into useful products using energy and other resources and learn about the ways in which chemical engineers make decisions and balance constraints to come up with new processes and products The author presents material and energy balances as tools to achieve a real goal workable economical and safe chemical processes and products Loaded with intriguing pedagogy this text is essential to a students first course in Chemical Engineering Additional resources intended to guide users are also available as package options such as ChemSkill Builder

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