

6. Write a pair of balanced half-reaction equations—one showing a gain of electrons and one showing a loss of electrons—for each of the following reactions:



9. For each of the following, write and label the oxidation and reduction half-reaction equations. Ignore spectator ions.



10. We have only looked at one type of single displacement reaction—a metal displacing another metal from an ionic compound. A nonmetal can also displace another nonmetal from an ionic compound. For example,



Using your knowledge of atoms and ions and the ideas presented in this chapter, write a pair of balanced half-reaction equations for this reaction—one showing a gain of electrons and one showing a loss of electrons.



Figure 9
The reduction of iron(III) oxide by aluminum is called the “thermite” reaction. Observe this reaction in

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