

Name: _____

AP Unit 6: Thermodynamics Designing a Hand Warmer Lab

Directions:

Objective: Determine the change in heat (q) and enthalpy (H) of a chemical process and determine the best design for a hand warmer.

Energy is essential for the functioning of societies and economies. For example, energy is needed to power the transportation sector, manufacturing, and the regulation of the temperature in homes. Hand warmers are simple devices which release heat to warm hands and feet through a chemical reaction. The design of many hand warmers consists of two compartments: one that contains water and another that contains a solid that is soluble in water. When the two components combine and the solid dissolves in the water, the hand warmer releases thermal energy to the surroundings. A device called a calorimeter is used to measure the amount of thermal energy (heat) released or absorbed by a system. The change in the amount of thermal energy is recorded as a temperature change.

Guiding Question:

How can calorimetry be used to determine the change in heat and enthalpy of a process and how can this information be used to determine the best design for a hand warmer?

Background

Thermochemistry is an area of chemistry that is concerned with how energy changes in chemical reactions. Heat (q) is a type of energy that brings about a change in temperature and, at times, may be involved in chemical processes. This change in temperature occurs when heat transfers between the system and its surroundings. If heat flows out of the system, the value of q will be negative; if heat is absorbed by the system, the value of q will be positive.

The first law of thermodynamics states that energy cannot be created or destroyed and is, therefore, conserved. Understanding this concept is necessary when studying energy changes. If there is energy or heat lost by a system, it must be gained by the surroundings in order for the energy to be conserved. Sometimes, the process of dissolving a solid in water may release or absorb heat. If the amount of energy required to break the crystal lattice of the solute and the attractions in the water exceeds the energy released during the solvation process, heat will be absorbed by the system from the surroundings, in an endothermic process. However, if the energy required to break apart the crystal lattice of the solute and the attractions of the water is less than the energy released during the solvation process, heat will be released by the system to the surroundings in an exothermic process.

Enthalpy (H) is a function of thermodynamics that represents the flow of heat in and out of a system. When heat flows into a system, the change in enthalpy (ΔH) is positive; when heat flows out of a system, the ΔH is negative. The change in enthalpy of a reaction can be studied and determined

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