

Determination of the Solubility-Product Constant for a Sparingly Soluble Salt

A. Preparation of a Calibration Curve

Initial $[\text{CrO}_4^{2-}]$ 0.0024M

Volume of 0.0024 M K_2CrO_4	Total volume	$[\text{CrO}_4^{2-}]$	Absorbance
1. <u>1 mL</u>	<u>50 mL</u>	<u>0.00048M</u>	<u>0.190</u>
2. <u>5 mL</u>	<u>50 mL</u>	<u>0.0024M</u>	<u>1.039</u>
3. <u>10 mL</u>	<u>50 mL</u>	<u>0.0048M</u>	<u>2.127</u>
4. <u>15 mL</u>	<u>50 mL</u>	<u>0.0072M</u>	<u>2.409</u>

Molar absorption coefficient for $[\text{CrO}_4^{2-}]$

1. _____ 2. _____ 3. _____ 4. _____

Average molar absorption coefficient _____

Standard deviation (show calculations) _____

B. Determination of the Solubility-Product Constant

Absorbance	$[\text{CrO}_4^{2-}]$	$[\text{Ag}^+]$	K_{sp} of Ag_2CrO_4
1. <u>0.603</u>	_____	_____	_____
2. <u>0.753</u>	_____	_____	_____
3. <u>1.286</u>	_____	_____	_____

Determination Of A Solubility Product Lab Answers

Julius Stieglitz



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