

Determination of K_a : Titration of a Weak Acid Lab Report

11/2/22

Introduction:

The goal of this lab is to determine the acid dissociation constant, K_a , by using Logger Pro and LabPro to follow the pH changes during the acid-base titration of the weak acid, CH_3COOH (acetic acid). The K_a values will come from the pH of the original solution and the pH of the half-neutralization point. We will also be finding the $[\text{HA}]$ concentration using the equivalence volume by looking at our titration curve.

Equilibrium is represented by the acid dissociation constant (K_a):

$K_a =$

The equation of the reaction is as follows:



Procedures:

- 1) Add 20.0 mL of weak acid solution into 100 mL beaker

Determination Of Ka Lab Report Answers

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A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the University of California, Berkeley text.

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