

Determination of Chloride using Potentiometry

1. Purpose

This procedure will determine the concentration of chloride ion with a chloride specific ion electrode using potentiometry.

2. Background

Potentiometry is an electrochemical method in which the potential of an electrochemical cell is measured while little to no current is passed through the sample. In titrimetric methods, this measurement can be used to indicate the end point. When an ion selective electrode is used, the measured potential is related to the ion concentration in solution and a quantitative determination can directly be made. The instrumentation used to perform potentiometry is straightforward and relatively inexpensive, consisting of an indicator electrode, a reference electrode, and a potential measuring device. One of the most common applications of potentiometry is in the measurement of pH.

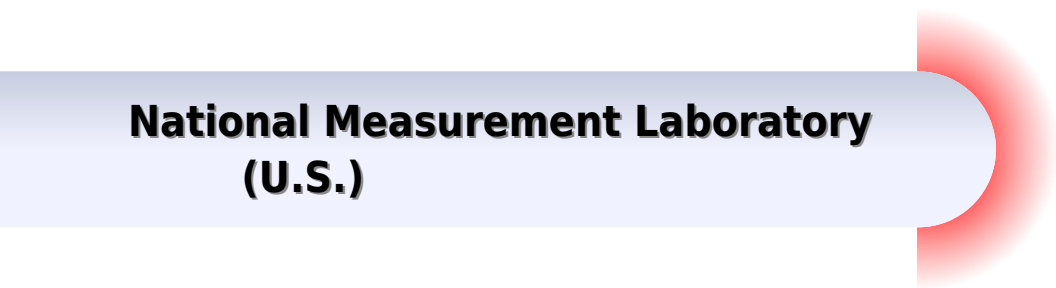
As indicated by the name, ion-selective electrodes (ISE's) possess a high degree of selectivity. These electrodes are routinely used in clinical laboratories to determine various ion concentrations (such as calcium ion) in blood samples. In this laboratory, the electrode used is specific for chloride ion. Because specific ion electrodes measure activity and not concentration, a large amount of an inert strong electrolyte (e.g. nitrate ion) can be added to fix the ionic strength to a constant value. When the ionic strength is constant, the activity is constant and concentration can be accurately measured. In this laboratory you will fix the ionic strength with an ionic strength adjustment buffer (ISAB). You will be asked to determine the concentration of chloride in an unknown sample at the ppm level.

The chloride ion selective electrode you will use is a crystalline solid-state electrode that contains a membrane, as shown in the diagram on the next page (Figure 1). The membrane consists of a solid salt of silver sulfide / silver chloride ($\text{Ag}_2\text{S} / \text{AgCl}$). The membrane must be insoluble in the analyte solution and contain the analyte ion of interest. The membrane is placed at the end of a solid plastic tube. This membrane is in contact with the analyte solution during the measurement. Inside of the tube is a reference solution, which contains a known and fixed concentration of analyte (Cl^-) solution. The concentration difference between this inner solution and analyte solution causes the migration of charged species across the membrane. This ion exchange process at the surface of the membrane causes a potential to develop. Since the potential of both the reference electrode and the inner reference (immersed in the standard solution) are constant, any change in measured potential is caused only by a change in potential across the membrane and is a function of the analyte chloride ion activity (or concentration).

It is also important to note that Figure 1 depicts a reference electrode external to the ISE (indicator electrode). However, a reference electrode can also be placed inside of the tube containing the ISE. This is called a "combination electrode" and is the type that you will be using in this laboratory experiment.

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Analysis of the Chloride Ion Concentration in Aqueous Solutions by Potentiometric Titration Lyndsie Selwyn, Canadian Conservation Institute, 2001-01-01 *Determination of Chloride by Precision Null-point Potentiometry*
James Dudley Winefordner, 1958 **DETERMINATION OF CHLORIDE BY PRECISION NULL-POINT POTENTIOMETRY**, 1959 **Determination of Chloride by Precision Null-point Potentiometry and Electrolytic Generation of Silver Ion** Gary Leroy Mowder, 1963 A Reference Method for the Determination of Chloride in Serum
National Measurement Laboratory (U.S.), R. A. Velapoldi, 1979 I. Automatic Differential Potentiometric Titrations. Ii. The Determination of Chloride in Titanium Sponge, 1955 **Automation of a Potentiometric Concentration Cell Apparatus and Its Use for the Determination of Chloride in Biological Fluids** Harley Harris Ross, 1957 **A Potentiometric Method for Determination of Sodium Chloride in Food** K. Ishijiki, 1978 **Potentiometry and Potentiometric Titrations** E. P. Serjeant, 1984-06-29 Cells electrodes and basic procedures rudiments of potentiometry classifications of potentiometry and titrimetry classifications and descriptions of electrodes procedures of analytical potentiometry electrodes for potentiometric titrimetry Applications to the determination of solution equilibrium data the determination of ionization constants in aqueous solution acid base equilibria in nonaqueous solvents the determination of stability constants by potentiometric titration Applications to titrimetric analysis the basics of titrimetry acid base titrimetry complexometric titrations oxidation reduction titrations precipitation titrations **Potentiometry and Ion Selective Electrodes (set Price of 34 Books)**
Alun Evans, 2008-09-23 The book provides a basic introduction to potentiometry and the use of ion selective electrodes in chemical analysis Written largely from first principles this self contained treatment develops the ideas behind the practice and explains the techniques of potentiometry in language the practitioner or student can understand The author explains procedures step by step and makes liberal use of worked examples and problems Thus the text is best worked through rather than just read or memorized Summary of Basic Principles General Principles of Potentiometry Method of Operation of Ion Selective Electrodes Applications **Potentiometric Micro-determination of Chloride in Sea Water ...** John Charles Illman, 1943 **Automatic Differential Potentiometric Titrations ; the Determination of Chloride in Titanium Sponge** E. Reinold Fett, 1955 **A Reference Method for the Determination of Chloride in Serum** National Measurement Laboratory (U.S.), 1979 **Potentiometric Water Analysis** Derek Midgley, Kenneth Torrance, 1978
Determination of Chloride in TATB. Progress Report, October--December 1975. [Potentiometric Analysis]., 1975 To determine total chloride content in TATB the sample was burned in a Parr Bomb and the chlorides formed were titrated potentiometrically using silver nitrate The electrode system is a chloride specific ion sensing element and a single junction reference Results on seven test samples agree generally within 10 percent with previous analyses by the Development Division Replicate analyses show precision of the method to be 10 relative percent Inorganic chloride is assayed by dissolving

the TATB in concentrated sulfuric acid The HCl which is liberated is swept into dilute base by a nitrogen sweep and analyzed potentiometrically with the chloride specific electrode Results on three test samples agree within 5 percent with Development Division analyses with a 2 relative percent precision of analysis

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web fig 2 picaxe 18m2 pictured above is the 18m2 chip based on the microchip pic16 l f1826 it has 2048 bytes of program memory flash ram 256 bytes of static ram 256 bytes of eeprom and 16 i o pins that can be configured as shown above the i o is configured in two 8 bit ports known as portb and portc

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by moving a servo to a predetermined position learn how to combine a picaxe ir sensors and a servo motor to create a device that responds to a moving stimulus spoilers i hook it up to eyeballs that can follow a finger

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web product details the smallest but most popular picaxe microcontroller very popular in education due to its low cost and small size but also very popular in a wide range of hobbyist projects small but perfectly formed this chip supports up to 6 inputs outputs with 3 analogue touch sensor channels

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web jan 5 2015 08m2 5 3 in or out 1 in only 1 out shared with programming pin 0 3 18m2 16 14 in or out 1 in only 1 out shared with programming pin 0 10 20m2 16 16 in or out 0 11 when it comes to selecting a chip size for a project it is important to make sure it has enough inputs and outputs to fulfil the specification of the project

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web apr 13 2013 the program code is simple and the parts count is insanely low at under 10 components total a picaxe 08m2 is in control of a small lcd counter module the code for the picaxe is also very simple only 11 lines long the picaxe takes care of 3 crucial functions it creates the 1 second 1000 msec sampling time such that the digital

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web step 1 picaxe project parts list 1 picaxe 08m2 microcontroller 1 8 ohm speaker or 40 ohm speaker and omit the 33 ohm resistor when using 40 ohms 1 8 pin ic socket 1 3mm stereo jack 1 22k ohm resistor 1 33 ohm resistor 3 10k ohm resistors 1 two position dip switch or equivalent switches