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Atomic Structure and the Periodic Table Study Guide

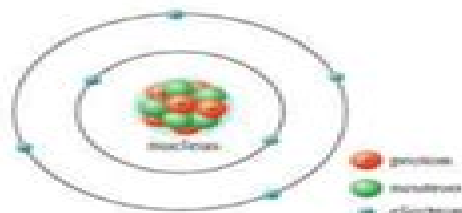
1. What is an atom?

Smallest particle of an element that retains the properties of that element.

2. Complete the following chart about subatomic particles.

Name	Charge	Symbol	Location	Relative Mass
Proton	+	p^+	Nucleus	1 amu
Electron	-	e^-	Electron cloud	0
neutron	0	n^0	nucleus	1 amu

3. Draw an atom and label the subatomic particles.

4. Which subatomic particle is used to determine the identity of an element?
protons5. Where is most of the mass of the atom? What particle(s) are in this area?
Nucleus; protons and neutrons

6. What is an atom called after it has gained or lost electrons? ___ion_____

7. What do you call an atom if it has more or less neutrons than normal? ___isotope_____

8. Label the information provided in the periodic table.

8	←	_____	Atomic number
O	←	_____	Symbol
Oxygen	←	_____	Name
15.999	←	_____	Atomic mass

9. From a box on the periodic table, how do you determine the number of protons, electrons, and neutrons?

1) # p^+ = _____ atomic number _____2) # e^- = _____ number of protons _____3) # n^0 = _____ atomic mass - atomic number _____

10. What does the atomic number represent? # protons in the nucleus

11. What does the atomic mass represent? # particles in the nucleus (protons & neutrons)

Atomic Structure And Periodic Relationships Study Guide

Kenneth W. Watkins



Atomic Structure And Periodic Relationships Study Guide:

Barron's Science 360: A Complete Study Guide to Chemistry with Online Practice Barron's Educational Series, Mark Kernion, Joseph A. Mascetta, 2021-09-07 provides a complete guide to the fundamentals of chemistry Page 4 of cover

CLEP Official Study Guide The College Board, 2005-08-20 The CLEP Official Study Guide 17th Edition is the only complete source of information about the CLEP exams You get Descriptions of the exams to help students decide which exams to take Sample questions and answers for all 35 exams Valuable test taking tips and information on obtaining credit through CLEP With more than 6 million CLEP exams taken since 1967 the College Level Examination Program is a credit by examination program that has proven to be effective for students advisors and faculty Students can demonstrate their comprehensive subject knowledge gain credit and move on to more advanced topics Advisors can help students plan their undergraduate years by helping them shorten the time needed to complete their degree Faculty can use their undergraduate introductory courses more effectively There are exams in 35 introductory college level subjects and are administered on computers at 1 300 college testing centers Approximately 2 900 colleges and universities grant credit for CLEP exams

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College of Engineering University of Michigan. College of Engineering,1970

Student's Guide to Fundamentals of Chemistry Jo A. Beran,2013-10-22

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Journal of Chemical Education ,1925 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific

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