

7&8 • Atomic Structure & Periodicity

PRACTICE TEST

$A = 2.18 \times 10^{-18} \text{ J}$	$h = 6.626 \times 10^{-34} \text{ J}\cdot\text{s}$
$R = 1.097 \times 10^7 \text{ m}^{-1}$	$c = 3.00 \times 10^8 \text{ m}\cdot\text{s}^{-1}$
mass of an electron = $9.11 \times 10^{-31} \text{ kg}$	

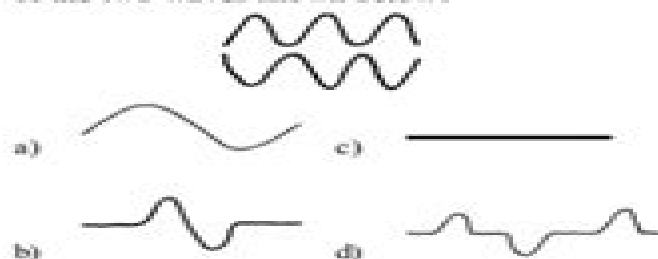
- What wavelength corresponds to a frequency of 8.22×10^9 Hz?
 - 0.307 m
 - 0.0365 m
 - 0.122 m
 - 0.110 m
 - 27.4 m
- A radio station transmits at 110 MHz (110×10^6 Hz). What wavelength is this radio wave?
 - 3.65×10^{-5} m
 - 3.30 m
 - 3.81×10^{-5} m
 - 2.73 m
- Which one of the following is NOT a proper unit for frequency?
 - Hz
 - s^{-1}
 - $m \cdot s^{-1}$
 - $\frac{1}{sec}$
- Calculate the wavelength of the fourth line in the Balmer series (the visible series) of the hydrogen spectrum.
 - 0.12334 m
 - 24.373 m
 - 4.1029×10^{-7} m
 - 36.559 m
 - 2.7353×10^{-7} m
- What is the relationship between the energy of a photon of light and its frequency?
 - $E = v$
 - $E = \frac{h}{v}$
 - $E = hv$
 - $E = \frac{1}{hv}$
 - $E = \frac{v}{h}$
- What is the energy needed to raise an electron in the hydrogen atom from the second energy level to the third energy level?
 - 1.52×10^{-4} J
 - 3.63×10^{-19} J
 - 2.18×10^{-19} J
 - 4.48×10^{-19} J
 - 3.03×10^{-19} J
- What is the de Broglie wavelength of an electron moving at 80.0% the speed of light.
 - 3.03×10^{-12} m
 - 2.42×10^{-12} m
 - 3.30×10^{11} m
 - 1.59×10^{-25} m
- What resultant is expected from the interference of the two waves shown below?

a)

b)

c)

d)
- Which quantum number determines the subshell occupied by an electron (s, p, d, f, etc.)?
 - n
 - l
 - m_l
 - m_s



Atomic Structure And Periodicity Practice Test Answers

**Barron's Educational Series, Neil D.
Jespersen, Pamela Kerrigan**



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