

8. As lone-pair electrons participate in
 - (a) in intermolecular interactions (hydrogen bonds and other secondary bonding/interactions)
 - (b) in surface (and adsorption) and catalytic reactions (adsorption on metal surfaces, etc.)
 - (c) in coordination and chemical bonding (coordination bonds) as well as in hydrogen bonds.
 Such systems can also happen to have an overall neutral or charged character that influences their
9. The order of increasing ΔE for electrons in shells, subshells, orbitals within the same subshell is:
 - (a) $ns < np < nd < nf$
 - (b) $ns < nd < np < nf$
 - (c) $ns < np < nd < nf$
 - (d) $ns < nd < nf < np$
10. The number of electrons present in complex of hexafluorophosphate ($[PF_6]^-$), perchlorate ($[ClO_4]^-$) are
 - (a) 48, 32
 - (b) 48, 36
 - (c) 48, 34
 - (d) 48, 38
11. Chemically, unreactive substances (e.g. noble gases) have ionization enthalpies that are comparatively
 - (a) very low
 - (b) very high
 - (c) moderate
 - (d) low
12. The ionization enthalpy (in eV) of a sample atom is found to be 13.6 eV for the ground state and 3.4 eV for
 - (a) $n=2$, $l=0$
 - (b) $n=2$, $l=1$
 - (c) $n=3$, $l=0$
 - (d) $n=3$, $l=1$
13. Elements with electronic configuration $ns^2(n-1)d^5$ are
 - (a) Mn, Cr, Fe
 - (b) Cr, Mn, Fe
 - (c) Mn, Cr, Fe
 - (d) Fe, Cr, Mn
14. Elements with electronic configuration $ns^2(n-1)d^5$ are
 - (a) Mn, Cr, Fe
 - (b) Cr, Mn, Fe
 - (c) Mn, Cr, Fe
 - (d) Fe, Cr, Mn
15. Which of the following correctly represents the electronic configuration of Chromium
 - (a) $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2$
 - (b) $1s^2 2s^2 2p^6 3s^2 3p^4 4s^1$
 - (c) $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2$
 - (d) $1s^2 2s^2 2p^6 3s^2 3p^4 4s^1$
16. Which of the following represents the correct order
 - (i) atomic number < ionization enthalpy < electronegativity < metallic character
 - (ii) electronegativity < ionization enthalpy < atomic number < metallic character
 - (iii) electronegativity < electronegativity < ionization enthalpy < metallic character
 - (iv) atomic number < ionization enthalpy < electronegativity < metallic character
 - (a) i, ii, iii, iv
 - (b) ii, iii, iv, i
 - (c) iii, iv, i, ii
 - (d) iv, iii, ii, i
17. Which of the following represents the correct order
 - (i) electronegativity < ionization enthalpy < atomic number < metallic character
 - (ii) electronegativity < ionization enthalpy < atomic number < metallic character
 - (iii) electronegativity < electronegativity < ionization enthalpy < metallic character
 - (iv) atomic number < ionization enthalpy < electronegativity < metallic character
 - (a) i, ii, iii, iv
 - (b) ii, iii, iv, i
 - (c) iii, iv, i, ii
 - (d) iv, iii, ii, i

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**Linda S. Cordell, Kent Lightfoot, Francis
McManamon, George Milner**

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