

I) Choose the correct answer

10 Marks

- To follow up saponification of ester with alkali by conductance (L), the following order is expected

- a) $L_0 < L_t < L_\infty$ b) $L_0 > L_t > L_\infty$ c) $L_0 < L_\infty < L_t$ d) $L_t > L_0 < L_\infty$ e) None

- In iodination of cyclohexanone, one observes that

- a) Protons are released b) A at $\lambda_{\max}$ decreases c) Color appears d) anions disappear

- Normality of 83.5 wt% H_2SO_4 with density = 1.442 g/cm^3 ($H=1$, $O=16$, $S=32$) is

- a) 18.1 b) 24.6 c) 10.5 d) 12.3 e) none

- A conductivity meter has a cell with two electrodes 2 cm^2 each and apart 0.1 cm , the cell constant is

- a) 2.0 cm^{-1} b) 1.0 cm^{-1} c) 10 cm^{-1} d) 20 cm^{-1} e) 0.05 cm^{-1}

- The 0.1 M salt solution with the highest ionic strength is

- a) KCl b) $AgNO_3$ c) $CaCl_2$ d) $Al_2(SO_4)_3$ e) $CuSO_4$

- 11.0% sucrose solution has a rotation angle of 13.5° . The sucrose percent at an angle of 37.2° is

- a) 20.9% b) 30.3% c) 27.6% d) 40.2% e) 36.3%

- For the iodination of cyclohexanone, one cannot follow the reaction kinetics by

- a) Spectrophotometer b) Polarimeter c) Potentiometer d) Thiosulphate titration

- During acid hydrolysis of sucrose the rotation angle of polarimeter moves anticlockwise because

- a) Sucrose is D b) Glucose is D c) Glucose/Fructose is weak L d) Both a & c e) All a-c

- The pH of a solution composed of 0.1 M acetic acid + 0.1 M sodium acetate ($K_a = 1.8 \times 10^{-5}$) is

- a) 4.0 b) 4.7 c) 5.0 d) 3.5 e) none

- The activity of K_2SO_4 ($a_{\pm K_2SO_4}$) in terms of the mean activity of the salt is given by

- a) $a_{\pm K_2SO_4}^2$ b) $a_{\pm K_2SO_4}^3$ c) $a_{\pm K_2SO_4}^4$ d) $a_{\pm K_2SO_4}^5$ e) $a_{\pm K_2SO_4}^6$

II) In a saponification reaction of 0.02 M ethyl acetate with 0.02 M KOH , the following conductance, L ,

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