

Physical Science Chapter 7 Study Guide

- **Molecules**—groups of two or more atoms linked by chemical bonds to form distinct units
- **Molecular mass**—the sum of the mass of the atoms in a molecule
- **Isomers**—molecules with the same number and types of atoms but in a different arrangement
- **Elements**—substances composed of a single type of atom
- **Compounds**—substances composed of different types of atoms bonded together
- **Mixtures**—different kinds of atoms that are physically mixed but not chemically united
 - **Solution**—a mixture in which the ingredients are uniformly mixed
 - **Solute**—that which is dissolved in a substance
 - **Solvent**—the substance into which something is dissolved
 - **Solubility**—the ability of one substance to dissolve in another substance
 - **Precipitate**—solid particles that come out of a solution
 - **Heterogeneous mixture**—substances which are mixed but not completely or uniformly mixed
 - **Colloid**—a mixture that contains tiny suspended clumps of particles
- **Chemical bonds**—the attraction between atoms making up a molecule
 - **Covalent**—a bond which results from a sharing of electrons
 - **Single covalent bonds** means sharing one pair of electrons; double means two pairs, etc.
 - **Lewis structure**—a diagram that shows all of the atoms in a molecule and the valence electrons around each atom
 - **Polar vs. nonpolar covalent bonds**
 - **Polar bond**—the one atom has a stronger pull on the electrons (this is called its electronegativity) and therefore “hogs” the electrons most of the time, creating a negative electrical force around it and leaving the other atom with a positive electrical force
 - **Nonpolar bond**—the atoms in the molecule have a basically equal pull on the electrons; therefore the charge remains balanced
 - **Ionic bond**—attraction between oppositely charged atoms; occurs when one atom transfers electrons to another atom
 - **Formula unit**—The simplest ratio of atoms in a nonmolecular compound
 - **Metallic bond**—occurs between metal elements and involves the sharing of electrons around many nuclei (think Hutterite colony)
- **Chemical formulas**
 - **Molecular formula**—shows the exact number and kind of atoms in a molecule
 - **Structural formula**—shows the general arrangement of atoms in a molecule
 - **Empirical formula**—shows only the simplest ratio of atoms in molecule

Physical Science Chapter 7 Study Guide Answers

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Table of Contents Physical Science Chapter 7 Study Guide Answers

1. Understanding the eBook Physical Science Chapter 7 Study Guide Answers
 - The Rise of Digital Reading Physical Science Chapter 7 Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Chapter 7 Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Chapter 7 Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Chapter 7 Study Guide Answers
 - Personalized Recommendations
 - Physical Science Chapter 7 Study Guide Answers User Reviews and Ratings
 - Physical Science Chapter 7 Study Guide Answers and Bestseller Lists
5. Accessing Physical Science Chapter 7 Study Guide Answers Free and Paid eBooks
 - Physical Science Chapter 7 Study Guide Answers Public Domain eBooks
 - Physical Science Chapter 7 Study Guide Answers eBook Subscription Services
 - Physical Science Chapter 7 Study Guide Answers Budget-Friendly Options

6. Navigating Physical Science Chapter 7 Study Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science Chapter 7 Study Guide Answers Compatibility with Devices
 - Physical Science Chapter 7 Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Chapter 7 Study Guide Answers
 - Highlighting and Note-Taking Physical Science Chapter 7 Study Guide Answers
 - Interactive Elements Physical Science Chapter 7 Study Guide Answers
8. Staying Engaged with Physical Science Chapter 7 Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Chapter 7 Study Guide Answers
9. Balancing eBooks and Physical Books Physical Science Chapter 7 Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Chapter 7 Study Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Chapter 7 Study Guide Answers
 - Setting Reading Goals Physical Science Chapter 7 Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Chapter 7 Study Guide Answers
 - Fact-Checking eBook Content of Physical Science Chapter 7 Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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