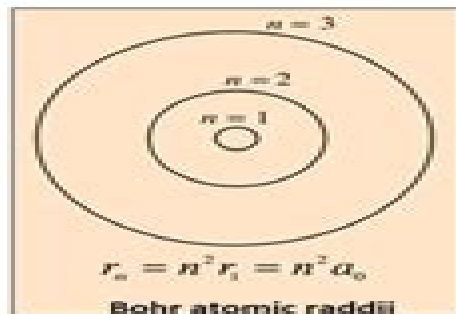




$$\Rightarrow r_2 = 2^2 \times 5.3 \times 10^{-11} \text{ m}$$

$$\Rightarrow r_2 = 2.12 \times 10^{-10} \text{ m}$$

For $n = 3$, we have,

$$\Rightarrow r_3 = (n)^2 r_1$$

$$\Rightarrow r_3 = 3^2 \times 5.3 \times 10^{-11} \text{ m}$$

$$\Rightarrow r_3 = 4.77 \times 10^{-10} \text{ m}$$

Hence, the radii of an electron for $n = 2$ and $n = 3$ orbits are $2.12 \times 10^{-10} \text{ m}$ and $4.77 \times 10^{-10} \text{ m}$.

9. A 12.5 eV electron beam is used to bombard gaseous hydrogen at room temperature. What series of wavelengths will be emitted?

9. It is given that the energy of the electron beam used to bombard gaseous hydrogen at room temperature is 12.5 eV. Also, the energy of the gaseous hydrogen in its ground state at room temperature is -13.6 eV.

When gaseous hydrogen is bombarded with an electron beam, the energy of the gaseous hydrogen becomes $-13.6 + 12.5 \text{ eV}$ i.e., -1.1 eV .

Orbital energy is related to orbit level (n) as:

$$E = \frac{-13.6}{(n)^2} \text{ eV}$$

$$\text{For } n = 3, E = \frac{-13.6}{9} = -1.5 \text{ eV}$$

This energy is approximately equal to the energy of gaseous hydrogen. It can be concluded that the electron has jumped from $n = 1$ to $n = 3$ level.

During its de-excitation, the electrons can jump from $n = 3$ to $n = 1$ directly, which forms a line of the Lyman series of the hydrogen spectrum.

We have the relation for wave number for Lyman series as:

$$\frac{1}{\lambda} = R_H \left(\frac{1}{1^2} - \frac{1}{n^2} \right)$$

Where,

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