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Chapter 1

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Answers to "Why?" in the text

Page 31:

Oxygen has an atomic mass of 16 whereas it is 14 for nitrogen. The O_2 molecule is therefore heavier than the N_2 molecule. Thus, from $\frac{1}{2}mv^2 = 3(\frac{1}{2}kT)$, the rms velocity of O_2 molecules is smaller than that of N_2 molecules.

Page 34, footnote 11

For small extensions, the difference between the engineering and instantaneous strains due to a temperature change are the same. Historically, mechanical and civil engineers measured extension by monitoring the change in length, ΔL ; and the instantaneous length L was not measured. It is not trivial to measure both the instantaneous length and the extension simultaneously. However, since we know L_o and measure ΔL , the instantaneous length $L = L_o + \Delta L$. Is the difference important? Consider a sample of length L_o that extends to a final length L due to a temperature change from T_o to T . Let $\epsilon = (L - L_o) / L_o = \Delta L / L_o$ be the engineering strain.

The engineering definition of strain and hence the thermal expansion coefficient is

$$\text{Engineering strain} = \frac{\delta L}{L_o} = \lambda \delta T$$

so that thermal expansion from T_o to T gives,

$$\int_{L_o}^L \frac{dL}{L_o} = \int_{T_o}^T \lambda dT \quad \therefore \quad \frac{\Delta L}{L_o} = \lambda(T - T_o) \quad \therefore \quad \epsilon = \lambda(T - T_o) \quad (1)$$

where $\epsilon = \Delta L / L_o$ is the engineering strain as defined above

Physics definition of strain and hence the thermal expansion coefficient is

$$\text{Instantaneous strain} = \frac{\delta L}{L} = \lambda \delta T$$

so that thermal expansion from T_o to T gives,

$$\int_{L_o}^L \frac{dL}{L} = \int_{T_o}^T \lambda dT \quad \therefore \quad \ln\left(\frac{L}{L_o}\right) = \lambda(T - T_o) \quad \therefore \quad \ln(1 + \epsilon) = \lambda(T - T_o) \quad (2)$$

We can expand the $\ln(1 + \epsilon)$ term for small ϵ , so that Equation (2) essentially becomes Equation (1)

1.1 Virial theorem The Li atom has a nucleus with a $+3e$ positive charge, which is surrounded by a full $1s$ shell with two electrons, and a single valence electron in the outer $2s$ subshell. The atomic radius of the

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1. Understanding the eBook Physics Solution Manual Chapter 14
 - The Rise of Digital Reading Physics Solution Manual Chapter 14
 - Advantages of eBooks Over Traditional Books
2. Identifying Physics Solution Manual Chapter 14
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physics Solution Manual Chapter 14
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physics Solution Manual Chapter 14
 - Personalized Recommendations
 - Physics Solution Manual Chapter 14 User Reviews and Ratings
 - Physics Solution Manual Chapter 14 and Bestseller Lists
5. Accessing Physics Solution Manual Chapter 14 Free and Paid eBooks
 - Physics Solution Manual Chapter 14 Public Domain eBooks
 - Physics Solution Manual Chapter 14 eBook Subscription Services

- Physics Solution Manual Chapter 14 Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Physics Solution Manual Chapter 14 Compatibility with Devices
 - Physics Solution Manual Chapter 14 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physics Solution Manual Chapter 14
 - Highlighting and Note-Taking Physics Solution Manual Chapter 14
 - Interactive Elements Physics Solution Manual Chapter 14
- 8. Staying Engaged with Physics Solution Manual Chapter 14
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physics Solution Manual Chapter 14
- 9. Balancing eBooks and Physical Books Physics Solution Manual Chapter 14
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physics Solution Manual Chapter 14
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physics Solution Manual Chapter 14
 - Setting Reading Goals Physics Solution Manual Chapter 14
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physics Solution Manual Chapter 14
 - Fact-Checking eBook Content of Physics Solution Manual Chapter 14
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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