

DALTON'S ATOMIC THEORY

Dalton's atomic theory was given by John Dalton in 1808. The main postulates of Dalton's atomic theory are:

1. Elements consist of indivisible, indestructible particles called atom.
2. Atoms of an element are identical to each other. They have the same mass and size.
3. Atoms of different elements differ in properties and have different masses and sizes.
4. Compounds are formed when atoms of different elements combine with each other in simple numerical ratio such as 1: 1, 2: 2, 2: 3 and so on.
5. Atoms cannot be created, destroyed or transformed into atoms of other elements.

The main failures of Dalton's atomic theory are:

1. It failed to explain how atoms of different elements differ from each other i.e., did not tell anything about structure of the atom.
2. It could not explain how and why atoms of different element combine with each other to form compound.
3. It failed to explain the nature of forces that bind together different atoms in a molecule.
4. It failed to explain Gay Lussac's law of combining volumes.
5. It did not make any distinction between ultimate particle of an element that takes part in reaction (atoms) and ultimate particle that has independent existence (molecules).

Prefixes Used With Units

The S.I. system recommends the multiples such as 10^3 , 10^6 , 10^9 etc. and fraction such as 10^{-3} , 10^{-6} , 10^{-9} etc., i.e. the powers are the multiples of 3. These are indicated by special prefixes. These along with some other fractions or multiples in common use, along with their prefixes are given below in Table and illustrated for length (m)

TABLE: SOME COMMONLY USED PREFIXES WITH THE UNITS

Prefix	Symbol	Multiplication Factor	Example
deci	d	10^{-1}	1 decimetre (dm) = 10^{-1} m
centi	c	10^{-2}	1 centimetre (cm) = 10^{-2} m
milli	m	10^{-3}	1 millimetre (mm) = 10^{-3} m
micro	μ	10^{-6}	1 micrometre (μ m) = 10^{-6} m
nano	n	10^{-9}	1 nanometre (nm) = 10^{-9} m
pico	p	10^{-12}	1 picometre (pm) = 10^{-12} m
femto	f	10^{-15}	1 femtometre (fm) = 10^{-15} m
atto	a	10^{-18}	1 attometre (am) = 10^{-18} m
zepto	z	10^{-21}	1 zeptometre (zm) = 10^{-21} m
yocto	y	10^{-24}	Yocto = 10^{-24} m
deca	da	10	1 dekametre (dam) = 10^1 m
hecto	h	10^2	1 hectometre (hm) = 10^2 m
kilo	k	10^3	1 kilometre (km) = 10^3 m
mega	M	10^6	1 megametre (Mm) = 10^6 m
giga	G	10^9	1 gigametre (Gm) = 10^9 m
tera	T	10^{12}	1 terametre (Tm) = 10^{12} m
peta	P	10^{15}	1 petametre (Pm) = 10^{15} m
exa	E	10^{18}	1 exametre (Em) = 10^{18} m
Zetta	Z	10^{21}	1 zetametre (Zm) = 10^{21} m
yotta	Y	10^{24}	Yotta = 10^{24} m

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**Timothy Scot Albrecht, California State
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Table of Contents Mole Concept Theory Notes Pdfslibforme

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 - The Rise of Digital Reading Mole Concept Theory Notes Pdfslibforme
 - Advantages of eBooks Over Traditional Books
2. Identifying Mole Concept Theory Notes Pdfslibforme
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mole Concept Theory Notes Pdfslibforme
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mole Concept Theory Notes Pdfslibforme
 - Personalized Recommendations
 - Mole Concept Theory Notes Pdfslibforme User Reviews and Ratings
 - Mole Concept Theory Notes Pdfslibforme and Bestseller Lists

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 - Mole Concept Theory Notes Pdfslibforme eBook Subscription Services
 - Mole Concept Theory Notes Pdfslibforme Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Mole Concept Theory Notes Pdfslibforme Compatibility with Devices
 - Mole Concept Theory Notes Pdfslibforme Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mole Concept Theory Notes Pdfslibforme
 - Highlighting and Note-Taking Mole Concept Theory Notes Pdfslibforme
 - Interactive Elements Mole Concept Theory Notes Pdfslibforme
 8. Staying Engaged with Mole Concept Theory Notes Pdfslibforme
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mole Concept Theory Notes Pdfslibforme
 9. Balancing eBooks and Physical Books Mole Concept Theory Notes Pdfslibforme
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mole Concept Theory Notes Pdfslibforme
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Mole Concept Theory Notes Pdfslibforme
 - Setting Reading Goals Mole Concept Theory Notes Pdfslibforme
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Mole Concept Theory Notes Pdfslibforme
 - Fact-Checking eBook Content of Mole Concept Theory Notes Pdfslibforme
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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