

KEY

35. Calculate the *maximum concentration* of fluoride ion possible in a solution in which $[\text{Sr}^{2+}] = 5.0 \times 10^{-3} \text{ M}$.



$$K_{\text{sp}} = [\text{Sr}^{2+}][\text{F}^{-}]^2$$

$$[\text{F}^{-}] = \sqrt{\frac{K_{\text{sp}}}{[\text{Sr}^{2+}]}} = \sqrt{\frac{4.3 \times 10^{-9}}{5.0 \times 10^{-3}}} = 9.3 \times 10^{-4} \text{ M} \quad \text{Answer } [\text{F}^{-}] = 9.3 \times 10^{-4} \text{ M}$$

36. Predict what would happen to the *solubility* of $\text{PbSO}_4(\text{s})$ if some K_2SO_4 solution is added. Explain your answer and include the use of an equilibrium equation.



When K_2SO_4 (soluble) is added $[\text{SO}_4^{2-}]$ increases causing a shift to the LEFT which decreases the solubility of PbSO_4 .

37. The solubility of ZnCO_3 in water is quite low. What could you add to increase the solubility?

Add an acid or Add a solution which precipitates Zn^{2+} or CO_3^{2-}
eg.) Na_2S

Explain fully how your method would work. Include the use of equilibrium equations.



$\text{H}^{+} \downarrow$ Addition of an acid will supply H^{+} , forming H_2O & CO_2 and decreasing the $[\text{CO}_3^{2-}]$. This causes a shift to the right which increases the solubility of ZnCO_3 .

$\text{H}_2\text{CO}_3 \rightarrow \text{H}_2\text{O} + \text{CO}_2(\text{g})$



$\text{S}^{2-} \downarrow$ Adding Na_2S will supply S^{2-} which will precipitate the Zn^{2+} , decreasing $[\text{Zn}^{2+}]$ thus causing a shift to the RIGHT which increases the solubility of $\text{ZnCO}_3(\text{s})$.

$\text{ZnS}(\text{s})$

Nelson Chemistry 12 Chapter 5 Solutions

Blake Robert Bichlmeir



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Liquid-State Physical Chemistry Gijsbertus de With, 2013-07-15 For many processes and applications in science and technology a basic knowledge of liquids and solutions is a must Gaining a better understanding of the behavior and properties of pure liquids and solutions will help to improve many processes and to advance research in many different areas This book provides a comprehensive self contained and integrated survey of this topic and is a must have for many chemists chemical engineers and material scientists ranging from newcomers in the field to more experienced researchers The author offers a clear well structured didactic approach and provides an overview of the most important types of liquids and solutions Special topics include chemical reactions surfaces and phase transitions Suitable both for introductory as well as intermediate level as more advanced parts are clearly marked Includes also problems and solutions *Electrochemistry in Nonaqueous Solutions* Kosuke Izutsu, 2002-05-06 Nonaqueous solutions are equally indispensable to electrochemistry Here Kosuke Izutsu brilliantly illustrates the numerous aspects of this fascinating topic whether the focus be on physicochemical processes or analytical methods The author discusses solvation and solvent effects emphasizing dynamic aspects important reactions including ionic and supercritical media as well as advanced techniques in polarography and voltammetry Throughout he effortlessly manages to provide a comprehensive overview while also presenting the very latest developments A number of example applications further enhance the practical value of this book and give it the feel of a reference work Written for both users and specialists this volume represents a wealth of vital information and belongs on every bookshelf

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