

Half Modern Chemistry Review

CHAPTER 9: STOICHIOMETRY

The following pages contain the book (but not all) of the information for the chapter 9 test.

Focus on this content, but make sure to review class notes, activities, assessments, questions, etc.

If you study this document and NOTHING else, you should at least be able to PASS the test.

*** Test items will be recall, application, and/or application of this content. ***

OUTCOMES

- Calculate molar percent or mole fraction/mass percent (M/C)
- Calculate the molar mass of a substance and percent (M/C)
- 9.1. Determine percent of water from balanced chemical equations. (T & P)
- 9.2. Perform stoichiometry calculations both by mole to mole, mole to gram, gram to mole, and gram to gram. (P & P)
- 9.3. Calculate stoichiometry problems and determine the limiting reagent, excess reagent, and amount of product produced. (P & P)

9.1: INTRODUCTION TO STOICHIOMETRY

a. Introductory

- stoichiometry: stoichiometry -- calculations involving the mass relationships of elements in compounds.
- molar stoichiometry -- calculations involving the mass relationships between reactants and products in a chemical reaction.
- molar ratio -- is a conversion factor that relates the amount, in moles, of any two substances involved in a chemical reaction.

b. Chapter Highlights

- Stoichiometry involves the mass relationships between reactants and products in a chemical reaction.
- Stoichiometry is based on the law of conservation of mass.
- Relating one substance to another requires recognizing the amount of each substance in moles.
- A molar ratio is the conversion factor that relates the amount in moles of any two substances in a chemical reaction. The molar ratio is derived from the balanced equation.
 - Example: $2\text{FeCl}_3 + 3\text{H}_2\text{O} \rightarrow \text{Fe}_2\text{O}_3 + 6\text{HCl}$ What is the ratio of iron to oxygen moles? 4:3 or 3:4
- Amount of a substance is measured in moles, and mass of a substance is determined using molar mass such as grams, kilograms, or milligrams.
- Mass and amount of substances are quantities, substances moles and grams are units.
- A balanced chemical equation is necessary to solve any stoichiometric problem.

9.2: MASS STOICHIOMETRIC CALCULATIONS

a. Chapter Highlights

- In an ideal stoichiometric calculation, the mass of the amount of any reactant or product can be calculated if the balanced chemical equation and the molar mass amount of any other reactant or product is known.

9.3: LIMITING REACTANTS AND PERCENTAGE YIELD

Modern Chemistry Chapter 9 Section 1 Review Answers

Norma K Dunlap, Donna M Huryn



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Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 **A Complete Introduction to Modern NMR Spectroscopy** Roger S. Macomber, 1997-12-23 Clear accessible coverage of modern NMR spectroscopy for students and professionals in many fields of science Nuclear magnetic resonance NMR spectroscopy has made quantum leaps in the last decade becoming a staple tool in such divergent fields as chemistry physics materials science biology and medicine That is why it is essential that scientists working in these areas be fully conversant with current NMR theory and practice This down to basics text offers a comprehensive up to date treatment of the fundamentals of NMR spectroscopy Using a straightforward approach that develops all concepts from a rudimentary level without using heavy mathematics it gives readers the knowledge they need to solve any molecular structure problem from a complete set of NMR data Topics are illustrated throughout with hundreds of figures and actual spectra Chapter end summaries and review problems with answers are included to help reinforce and test understanding of key material From NMR studies of biologically important molecules to magnetic resonance imaging this book serves as an excellent all around primer on NMR spectroscopic analysis

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