

KINETICS Practice Problems and Solutions

Name: _____
Period: _____ Class: _____

AP Chemistry
Dr. Menden

The following questions represent generalized types of rate questions. Please answer each question completely and thoroughly. The solutions will be posted on Blackboard.

1. Please consider the reaction of H_2 and I_2 at 400°C.

- Write the balanced chemical equation for this reaction. **$\text{H}_2 + \text{I}_2 \rightarrow 2\text{HI}$**
- Write the rate law for this reaction. **$\text{rate} = k[\text{H}_2][\text{I}_2]$**
- Calculate the rate constant, k , for this reaction. **$k = 0.0012 \text{ M}^{-1}\text{s}^{-1}$**

2. Consider the reaction: $\text{P}_4 + 6\text{H}_2 \rightarrow 4\text{PH}_3$. It was found that the reaction was second order with respect to P_4 .

A rate study of this reaction was conducted at 240°C. The data that were obtained are shown in the table.

$[\text{P}_4]$, mol/L	$[\text{H}_2]$, mol/L	Initial Rate, mol/L·s
0.0010	0.0010	1.20×10^{-4}
0.0010	0.0020	1.20×10^{-4}
0.0020	0.0020	4.80×10^{-4}

- Write the rate law expression: **$\text{rate} = k[\text{P}_4]^2$**
- Write the rate law for this reaction: **$\text{rate} = 0.00012 [\text{P}_4]^2$**
- Calculate the value and units of the rate constant, k . **$1.2 \times 10^{-4} \text{ L}^2 \text{ mol}^{-2} \text{ s}^{-1}$**

3. Consider the reaction: $\text{NO}_2 + \text{CO} \rightarrow \text{NO} + \text{CO}_2$. It was found that the reaction was second order in NO_2 . The data that were obtained are shown in the table.

$[\text{NO}_2]$, mol/L	$[\text{CO}]$, mol/L	Initial Rate, mol/L·s
0.020	0.010	0.0010
0.020	0.020	0.0010
0.040	0.020	0.0040

- Write the rate law expression: **$\text{rate} = k[\text{NO}_2]^2$**
- Write the rate law for this reaction: **$\text{rate} = 0.00125 [\text{NO}_2]^2$**
- Calculate the value and units of the rate constant, k . **$1.25 \times 10^{-2} \text{ L}^2 \text{ mol}^{-2} \text{ s}^{-1}$**

4. Consider the following reactions: $\text{A}_2 \rightarrow \text{B}_2 \rightarrow \text{C}$ (slow) **$\text{A}_2 \rightarrow \text{B}_2 \rightarrow \text{C}$** (fast)

- Write the overall balanced chemical equation: **$\text{A}_2 \rightarrow \text{C}$**
- Write the rate law for this reaction: **$\text{rate} = k[\text{A}_2]$**
- What is the order with respect to each reactant? **A_2 is first order.**

Kinetics Problems And Solutions

Gorimir Gorimirovich Chernyĭ



Kinetics Problems And Solutions:

Kinetics of Metal Ion Adsorption from Aqueous Solutions Sotira Yiaccoumi, Chi Tien, 2013-11-27 This monograph is intended to provide a systematic presentation of theories concerning the adsorption of metal ions from aqueous solutions onto surfaces of natural and synthetic substances and to outline methods and procedures to estimate the extent and progress of adsorption. As heavy metals and the problems associated with their transport and distribution are of serious concern to human health and the environment, the materials presented in this volume have both theoretical and practical significance. In writing this monograph, one of our goals was to prepare a book useful to environmental workers and practicing engineers. For this reason, our presentation relies heavily on concepts commonly used in the environmental engineering literature. In fact, the volume was prepared for readers with a basic understanding of environmental engineering principles and some knowledge of adsorption processes. No prior familiarity with the ionic solute adsorption at solid solution interfaces is assumed. Instead, introduction of the necessary background information was included. Generally speaking, metal ion adsorption may be studied in terms of three distinct but interrelated phenomena: surface ionization, complex formation, and the formation and presence of an electrostatic double layer adjacent to adsorbent surfaces. Analyses of these phenomena with various degrees of sophistication are presented. xviii ADSORPTION OF METAL IONS FROM AQUEOUS SOLUTIONS presented and their various combinations yield different models that describe metal ion adsorption. *Problems in Chemical Kinetics* Nikolai Markovich Emanuel', 1981

Kinetics in Materials Science and Engineering Dennis W. Readey, 2017-01-27 A pedagogical gem. Professor Readey replaces black box explanations with detailed insightful derivations. A wealth of practical application examples and exercise problems complement the exhaustive coverage of kinetics for all material classes. Prof. Rainer Hebert, University of Connecticut. Prof. Readey gives a grand tour of the kinetics of materials suitable for experimentalists and modellers. In an easy-to-read and entertaining style, this book leads the reader to a fundamental model-based understanding of kinetic processes critical to development, fabrication, and application of commercially important soft polymers, biomaterials, hard ceramics, metals, and composite materials. It is a must-have for anyone who really wants to understand how to make materials and how they will behave in service. Prof. Bill Lee, Imperial College London. Fellow of the Royal Academy of Engineering. A much-needed text filling the gap between an introductory course in materials science and advanced materials-specific kinetics courses. Ideal for the undergraduate interested in an in-depth study of kinetics in materials. Prof. Mark E. Eberhart, Colorado School of Mines. This book provides an in-depth introduction to the most important kinetic concepts in materials science, engineering, and processing. All types of materials are addressed, including metals, ceramics, polymers, electronic materials, biomaterials, and composites. The expert author, with decades of teaching and practical experience, gives a lively and accessible overview explaining the principles that determine how long it takes to change material properties and make new and better materials. The chapters cover a broad range of topics extending from the heat treatment of steels to the

processing of silicon integrated microchips and the production of cement to the movement of drugs through the human body The author explicitly avoids black box equations providing derivations with clear explanations

Chemical Kinetics with Mathcad and Maple Viktor Korobov, Valery Ochkov, 2011-05-26 The authors explain at length the principles of chemical kinetics and approaches to computerized calculations in modern software suites mathcad and maple Mathematics is crucial in determining correlations in chemical processes and requires various numerical approaches Often significant issues with mathematical formalizations of chemical problems arise and many kinetic problems can't be solved without computers Numerous problems encountered in solving kinetics calculations with detailed descriptions of the numerical tools are given Special attention is given to electrochemical reactions which fills a gap in existing texts not covering this topic in detail The material demonstrates how these suites provide quick and precise behavior predictions for a system over time for postulated mechanisms Examples i.e oscillating and non isothermal reactions help explain the use of mathcad more efficiently Also included are the results of authors own research toward effective computations

The Application of metabolic and excretion kinetics to problems of industrial toxicology Jerzy Piotrowski, 1971

The Kinetic Theory of Gases Oskar Emil Meyer, 1899

Analytical Mechanics for Engineers Fred B. Seely, Newton Edward Ensign, 1921

Solutions and Solubilities Michael R. J. Dack, 1975

Physical and Chemical Processes in Gas Dynamics Gorimir Gorimirovich Chernyĭ, 2004 CD ROM CARAT software Computer interactive models and database for the Calculation of RATE constants Supplement to Reference Book Physical and chemical processes in gas dynamics Cross sections and rate constants

Thermal, Power and Electrical Engineering Charles C. Zhou, Guo Jun Ma, Ru Tang Liao, Jian Wen Wang, 2013-08-16

Selected peer reviewed papers from the 2013 2nd International Conference on Energy and Environmental Protection ICEEP 2013 April 19 21 2013 Guilin China

Calculations in Hydraulic Engineering: Fluid pressure, and the calculations of its effects in engineering structures Thomas Claxton Fidler, 1898

College of Engineering Catalogue Carnegie Institute of Technology. College of Engineering and Science, 1927

Nuclear Science Abstracts, 1971 NSA is a comprehensive collection of international nuclear science and technology literature for the period 1948 through 1976 pre dating the prestigious INIS database which began in 1970 NSA existed as a printed product Volumes 1 33 initially created by DOE's predecessor the U S Atomic Energy Commission AEC NSA includes citations to scientific and technical reports from the AEC the U S Energy Research and Development Administration and its contractors plus other agencies and international organizations universities and industrial and research organizations References to books conference proceedings papers patents dissertations engineering drawings and journal articles from worldwide sources are also included Abstracts and full text are provided if available

Handbook of Heat and Mass Transfer: Catalysis, kinetics, and reactor engineering, 1989

Introductory Nuclear Reactor Dynamics Karl Otto Ott, Robert J. Neuhold, 1985 This text presents the theory and methods of prediction that are the heart of nuclear reactor safety Time dependent reactor behavior is explained in both

mathematical and physical terms This book also explains the logic behind the working formulas and calculational methods for reactor transients and illustrates typical dynamic responses The classical concept of point kinetics is developed in three steps with discussion of various solutions to kinetics problems Each chapter includes homework problems and review questions

The Principles of chemistry v. 2 Dmitry Ivanovich Mendeleyev,1897

Chemical and Engineering News ,1958

Transactions of the American Nuclear Society American Nuclear Society,1992

Techniques of Chemistry ,1975

Paper ,1985

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