

Unit 5 Types of Chemical Reactions Lab Report (37 pts)

In this experiment, you will mix chemicals and observe synthesis, single displacement, and double displacement reactions. You will make observations about what happens when you combine the different chemicals. Have your solubility chart. You will write balanced chemical equations for each reaction. You will need screen shots of your results at the end so be sure take snips of your results and paste them in the end right from the start.

This is in the content at Investigation: Types of Chemical Reactions 1 and you have to go to slide two to get to it:

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Investigation: Types of Chemical Reactions 1 ~

(3 points)

Part 1 Synthesis combustion reaction:

- You burned a Magnesium (Mg(s)) strip in the presence of Oxygen ($\text{O}_2(\text{g})$) gas. Mg^{2+} , O^{2-}
 - What happened, describe what you saw, what happened in terms of left-over product, and add snip.
 - Write a **balanced** synthesis reaction for this process. I should see reactants < products
 - Using the A, B symbols, write a general equation for a synthesis reaction.

(4 points)

Part 2: Single-Displacement Reactions:

- For each of the four single-displacement reactions, describe what happened in each beaker. Use the Reactivity chart on page two to confirm your reaction.

Table 1 Observations Single Displacement Reactions			
Beaker	Solids (s)	Solutions (aq) 10 drops	Observations Before & After can add a screen shot
A	Aluminum	Copper sulfate $\text{CuSO}_4(\text{aq})$	Darker specks develop on aluminum, this is copper
B	Zinc	Copper sulfate $\text{CuSO}_4(\text{aq})$	
C	Zinc	Silver nitrate $\text{AgNO}_3(\text{aq})$	
D	Copper	Silver nitrate $\text{AgNO}_3(\text{aq})$	

- (1 pt) Using the A, B and C symbols, write a general equation for a single-displacement reaction.

(6 pts for products, balanced and parenthesis for solid, liquid, gas or aqueous)

- Write a **balanced** equation each reaction.

Ionic charges: Zn^{2+} , Al^{3+} , Cu^{2+} , Ag^{+} , SO_4^{2-} , NO_3^{-}

Here are the chemical formulas for the reactants in each reaction: I should see reactants < products. The reactants are given, but you need to cross down your oxidation numbers to get the correct chemical formulas for the products. Be sure to include (aq), or (s) for each chemical formula. I'll do the reaction A, you do the rest.

- A. Aluminum – Al(s) and copper sulfate – $\text{CuSO}_4(\text{aq})$

1. If the charges of the ions are the same, then the ratio is 1:1.
2. If the charges of the ions are different, then the ratio can be determined by criss-crossing the charges.



Lab Report Types Of Chemical Reactions Answers

Ralph Thomas Myers



Lab Report Types Of Chemical Reactions Answers:

STEM Road Map 2.0 Carla C. Johnson, Erin E. Peters-Burton, Tamara J. Moore, 2021-03-10 Featuring a team of over thirty STEM education professionals from across the United States the updated and revised edition of this landmark book provides an integrated STEM curriculum encompassing the entire K 12 spectrum with complete grade level learning based on a spiraled approach to building conceptual understanding Taking into account the last five years of evolution in STEM education the second edition includes an increased focus on computer science computational thinking mathematics and the arts as well as cultural relevance and addressing the needs of diverse learners and underrepresented students Divided into three main parts Conceptualizing STEM STEM Curriculum Maps and Building Capacity for STEM each section is designed to build common understandings of integrated STEM provide rich curriculum maps for implementing integrated STEM at the classroom level and offer supports to enable systemic transformation to an integrated STEM approach Written for teachers policymakers and administrators this second edition is fully updated to account for the needs of K 12 learners in the innovation age STEM Road Map 2 0 enables educators to implement integrated STEM learning into their classroom without the need for extensive resources empowering educators and supporting students

Experimental Organic Chemistry

Daniel R. Palleros, 2000-02-04 This cutting edge lab manual takes a multiscale approach presenting both micro semi micro and macroscale techniques The manual is easy to navigate with all relevant techniques found as they are needed Cutting edge subjects such as HPLC bioorganic chemistry multistep synthesis and more are presented in a clear and engaging fashion

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Holt Chemistry Ralph Thomas Myers, 2004

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Inquiry-based Experiments in Chemistry Valerie Ludwig Lechtanski, 2000 Inquiry Based Experiments in Chemistry is an alternative to those cookbook style lab manuals providing a more accurate and realistic experience of scientific investigation and thought for the high school chemistry or physical science student

An Introductory Guide to EC Competition Law and Practice Valentine Korah, 1994

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Government Reports Announcements & Index, 1988

Abstracts of Bacteriology Society of American Bacteriologists, 1924 Includes Scientific proceedings of the Society of American Bacteriologists

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General, Organic, and Biological Chemistry Study Guide and Selected Solutions Karen C. Timberlake, 2001-11 Keyed to the learning goals in the text this guide is designed to promote active learning through a variety of exercises with answers and mastery exams The guide also contains complete solutions to odd numbered problems

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