

Name _____ Class _____ Date _____

Lab – Shapes of Covalent Molecules

Introduction

The type of chemical bond that will form between two atoms can be predicted by calculating the difference in the atoms' **electronegativities**. When the values of two atoms' electronegativities are far apart, one atom loses one or more electrons to the other and an **ionic bond** is formed. When two atoms' electronegativities are closer in value, the atoms share pairs of electrons and a **covalent bond** is formed. Covalent bonds are considered **nonpolar** when the atoms' electronegativities are nearly the same and the electrons are shared evenly between the atoms. Covalent bonds are considered **polar** when the two atoms unequally share electron pairs.

Electron pairs around atoms, whether they are bonding or nonbonding (lone) pairs, all repel each other and try to move as far from each other as possible. According to **VSEPR theory** (valence shell electron pair repulsion theory), this arrangement of valence electron pairs determines the shape of covalent molecules.

When a covalent molecule contains only nonpolar bonds, the molecule as a whole is also nonpolar. Its electron cloud is evenly and symmetrically spread around the molecule so that there are no areas of partially positive or negative charge. When a covalent molecule contains polar bonds and the electron cloud is not symmetric, the molecule as a whole is polar and has areas of partial charge (dipoles). The molecule is partially negative (δ^-) near the more electronegative atom and partially positive (δ^+) near the less electronegative atom. In some cases, covalent molecules containing polar bonds are, as a whole, nonpolar if the "pulls" are symmetric.

Procedure

1. Construct the electron dot (Lewis) structure for the molecule given.
2. Determine the shape of the molecule by constructing the molecule.
3. Record the electronegativities of the elements and calculate the difference.
4. State the type of bond (polar covalent or nonpolar covalent) formed using the data provided by the teacher.
5. State whether the molecule as a whole is polar or nonpolar by asking yourself, "If there were a tug of war in this molecule, would anyone win?" If there is a winner, the molecule is polar, if there would be a tie, the molecule is nonpolar.

Formula	Electron dot structure	Shape of molecule	Calculate electronegativity difference	Bond (polar or nonpolar)	Molecule (polar or nonpolar)
H ₂					
F ₂					

Lab Shapes Of Covalent Molecules Answer Key

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Lab Shapes Of Covalent Molecules Answer Key:

Biological Investigations Lab Manual Warren Dolphin, David Vleck, Linda Westgate, James Colbert, 2010-01-27 The lead author of eight successful previous editions has brought together a team that combined has well over 60 years experience in offering beginning biology labs to several thousand students each year at Iowa State University Their experience and diverse backgrounds ensure that this extensively revised edition will meet the needs of a new generation of students Designed to be used with all majors level general biology textbooks the included labs are investigative using both discovery and hypothesis based science methods Students experimentally investigate topics observe structure use critical thinking skills to predict and test ideas and engage in hands on learning Students are often asked what evidence do you have that in order to encourage them to think for themselves By emphasizing investigative quantitative and comparative approaches to the topics the authors continually emphasize how the biological sciences are integrative yet unique An instructor s manual available through McGraw Hill Lab Central provides detailed advice based on the authors experience on how to prepare materials for each lab teachings tips and lesson plans and questions that can be used in quizzes and practical exams This manual is an excellent choice for colleges and universities that want their students to experience the breadth of modern biology **Holt**

Chemistry Ralph Thomas Myers, 2004 *Te HS&T a* Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2004-02

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