

Chapter Test B

Thermodynamics

MULTIPLE CHOICE

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. When an ideal gas does positive work on its surroundings, which of the gas's quantities increases?
- a. temperature
 - b. volume
 - c. pressure
 - d. internal energy
- _____ 2. Air cools as it escapes from a diver's compressed air tank. What kind of process is this?
- a. isovolumetric
 - b. isobaric
 - c. adiabatic
 - d. isothermal
- _____ 3. In an isovolumetric process for an ideal gas, the system's change in the energy as heat is equivalent to a change in which of the following?
- a. temperature
 - b. volume
 - c. pressure
 - d. internal energy
- _____ 4. How is conservation of internal energy expressed for a system during an adiabatic process?
- a. $Q = W = 0$, so $\Delta U = 0$ and $U_i = U_f$
 - b. $Q = 0$, so $\Delta U = -W$
 - c. $\Delta T = 0$, so $\Delta U = 0$; therefore, $\Delta U = Q - W = 0$, or $Q = W$
 - d. $\Delta V = 0$, so $P\Delta V = 0$ and $W = 0$; therefore, $\Delta U = Q$
- _____ 5. An ideal gas system undergoes an isovolumetric process in which 20 J of energy is added as heat to the gas. What is the change in the system's internal energy?
- a. -20 J
 - b. 0 J
 - c. 5 J
 - d. 20 J
- _____ 6. Which of the following is *not* a way in which a cyclic process resembles an isothermal process?
- a. Energy can be transferred as work.
 - b. Energy can be transferred as heat.
 - c. The temperature of the system remains constant throughout the process.
 - d. There is no net change in the internal energy of the system.

Assessment Chapter Test B

**Kathy Bonewit-West, Sue Hunt, Edith
Applegate**



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