

Chapter Test B

Chapter: States of Matter

Match each term with its definition, using the number of the term or phrase that best describes each statement or term. Write answers in the spaces.

1. According to the kinetic-molecular theory, particles of a substance
a. move about with both fast and slow motion.
b. move about with equal velocity.
c. move both the same amount and the same.
d. do not exhibit random but constant movement.
2. According to the equation $KE = \frac{1}{2}mv^2$, all particles within a substance
exhibit random motion and the same temperature. Particles within a substance
a. have different energy.
b. have equal energy.
c. exhibit average energy.
d. have average energy.
3. Liquids differ from solids mainly because they
a. expand when the temperature rises.
b. are composed of different particles than solids.
c. have particles in constant motion.
d. expand when pressure or temperature changes.
4. When heated, a gas expands and will
a. probably become liquid or solid.
b. move from a high temperature region.
c. exhibit a change in density resulting in expansion.
d. move in a constant slightly slower equilibrium temperature.
5. When heated, air flows in upward by a rapid upward movement of individual
air molecules and will
a. exhibit greater weight and expand.
b. exhibit greater weight and contract.
c. expand uniformly in upward and upward.
d. expand.
6. A solid substance shows random motion of particles
a. for a limited number of particles.
b. for a limited number of particles.
c. for a given number of particles.
d. for all particles.

Assessment Chapter Test B States Of Matter

Michael Hoelscher



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