

Chapter 10: General Organic Mechanisms

Structure of Test 2:

Name: _____

(Instructions, Readability)

1) The boiling points of **alkynes** hydrocarbons are determined by _____?

- A) van der Waals attractions
- B) ionic bonding
- C) hydrogen bonding
- D) dipole-dipole attractions
- E) **covalent bonding**

2) Water has a high boiling point because it is _____?

- A) less dense than water
- B) volatile and light in weight
- C) **hydrogen bond network**
- D) polar and water is nonpolar
- E) less dense than water

3) Draw out the following molecules or write the chemical equation showing how the substance dissociates and/or write the K_{sp} expression.

- BaSO_4
- $\text{Ca}_3(\text{PO}_4)_2$
- Ba_3N_2
- Ag_2S
- Ba_3N_2

Answer: $\text{BaSO}_4 \rightarrow \text{Ba}^{2+} + \text{SO}_4^{2-}$ $K_{sp} = [\text{Ba}^{2+}][\text{SO}_4^{2-}]$

Answer: $\text{Ca}_3(\text{PO}_4)_2 \rightarrow 3\text{Ca}^{2+} + 2\text{PO}_4^{3-}$ $K_{sp} = [\text{Ca}^{2+}]^3[\text{PO}_4^{3-}]^2$

Answer: $\text{Ba}_3\text{N}_2 \rightarrow 3\text{Ba}^{2+} + \text{N}_2^{4-}$ $K_{sp} = [\text{Ba}^{2+}]^3[\text{N}_2^{4-}]$

Answer: $\text{Ag}_2\text{S} \rightarrow 2\text{Ag}^{+} + \text{S}^{2-}$ $K_{sp} = [\text{Ag}^{+}]^2[\text{S}^{2-}]$

Answer: $\text{Ba}_3\text{N}_2 \rightarrow 3\text{Ba}^{2+} + \text{N}_2^{4-}$ $K_{sp} = [\text{Ba}^{2+}]^3[\text{N}_2^{4-}]$

General Organic And Biochemistry Lab Answers

**Frederick A. Bettelheim, Joseph Marvin
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